Elsevier Editorial System(tm) for Journal of Solid State Chemistry

Manuscript Draft

Manuscript Number: JSSC-19-979R1

Title: Synthesis and Optical Properties RE2O2S:Ln (RE = La, Y; Ln = Ce, Eu, Dy, Er)

Article Type: Research Paper

Keywords: oxysulfides, sulfates, rare-earth, reduction, optical ceramic, luminescence

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Abstract: The phase formation sequence was studied in the preparation of solid solutions of RE2O2S: Ln' (RE = La, Y; Ln' = Ce, Eu, Dy, Er) by the reduction of the match co-precipitated sulfates, followed by sulfidization of the reduction products. For uniform distribution of cations in the matrix, a method of chemical homogenization was used, consisting in the preparation of an aqueous solution containing all the necessary cations and their subsequent precipitation in the form of sulphates. The use of sulfates as precursors facilitates the process of obtaining solid solutions of oxysulfides, since sulfates already contain SO42- ions. The phase and morphological certification of the obtained solid solutions was carried out. The study of steady state luminescent properties demonstrated characteristic bands which are assigned to 4f-4f and 5d-4f transition. The obtained results showed the possibility of applying the method to synthesize optical ceramics based on solid solutionsRE2O2S: Ln (RE = La, Y; Ln = Ce, Eu, Dy, Er).

Elena I. Sal’nikova

Institute of Chemistry

Dear Anthony Powell

Our team of authors is grateful for the submitted reviews.
All comments and suggestions have been taken into account, as far as possible, which is reflected in the "Responses to reviewers" and the corrected version of the manuscript.
We confirm that this study presents original research.
We have no conflict of interest in this work.
We look forward to a long and fruitful cooperation with your magazine.

E.I. Salnikova

17/09/2019

Response to Reviews

The authors are grateful to the editor and reviewers for their comments on our article. Sy tried to take them all into account and make the necessary adjustments to the article. Below are our answers.

Reviewer #1: The article titled "Synthesis and Optical Properties RE2O2S:Ln (RE = La, Y; Ln = Ce, Eu, Dy, Er)" describes the preparation of activated oxysulfides and their properties. Although the article is short, authors have discussed the synthesis and characterizations in detail. The following questions might be helpful:

-How did authors certify the phase compositions of each materials? Is there any elemental analysis performed?

In this study, the phase composition was studied only by X-ray phase analysis, using the Rietveld method to process the results. The statistical processing parameters have good values for both multiphase and single-phase samples. Changing the parameters of an elementary cell completely correlates with the degree of doping, which confirms the formation of solid solutions, rather than two-phase samples.

-Thermogravimetric analysis/any other relevant experiment should be reported that supports the suggested reaction equations (e.g., page 5 line 35-38, etc.).

Unfortunately, we have no opportunity to conduct thermogravimetric analysis in a restorative atmosphere. The equations are based on the analysis of a large amount of experimental data. Solid reaction products are described based on x-ray phase analysis data. By-products are determined based on experimental observations in the synthesis process. Elemental sulfur and water are released at the very first stages of synthesis and condense on the cold walls of the reactor (Fig. 1). That allows you to uniquely identify them.

At certain stages of the synthesis at low temperatures, the off-gas does not give a reaction when sparging them through a solution of lead nitrate. However, when they are sparged through a barium nitrate solution, a white precipitate soluble in nitric acid precipitates. Based on this, it was concluded that at certain stages of the synthesis, a shallow reduction to sulfur dioxide can occur.

Unambiguous writing of the reaction equations is most likely impossible. All the written reaction equations are possible from the thermodynamic point of view and are most likely coupled during the process of sulfate reduction.

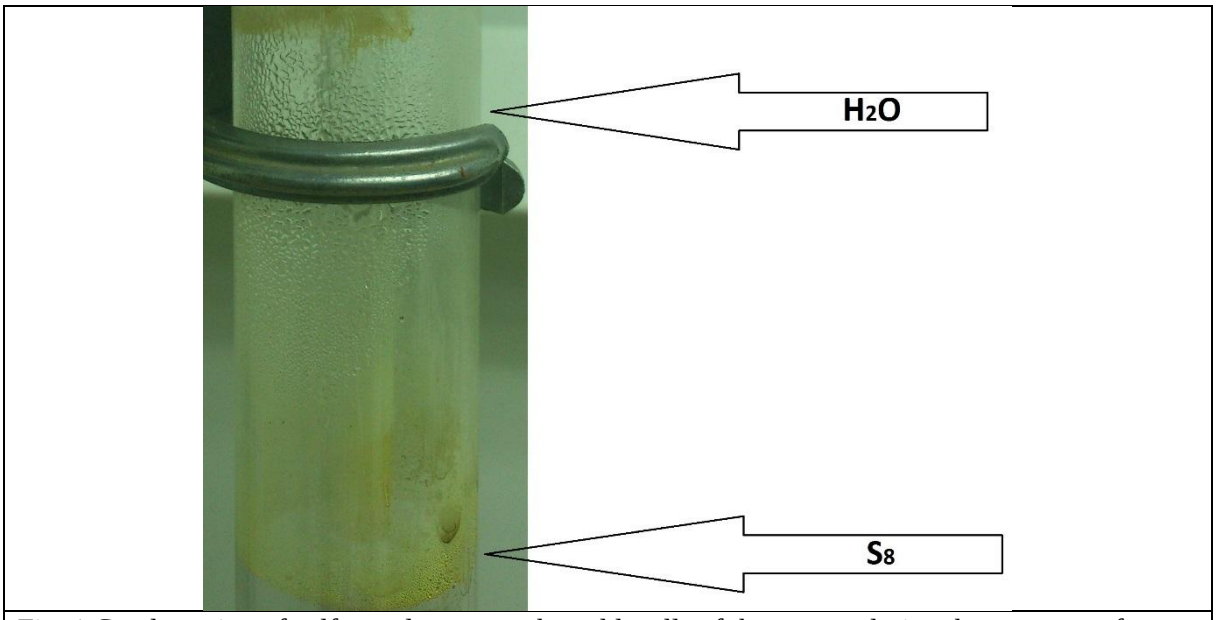

Fig. 1 Condensation of sulfur and water on the cold walls of the reactor during the recovery of sulfates.

-Graphical abstract for this article is missing.

Graphical abstract corrected and submitted

-It would be better to have sub-sections under results and discussions to explain different observations.

The remark is taken into account in the manuscript of the article.

-The discussion on morphological transformation requires more explanation and details (page 6, line 4-8, fig 2).

Corresponding discussions and comparison with the literature data are presented in the text of the article.

The article needs to address the above-mentioned concerns/questions before it can be considered in JSSC.

Reviewer #2: Manuscript Number JSSC-19-979 reports on a synthesis method for lanthanide doped sulfide oxides and their optical properties. There are some minor issues to be resolved before it can be published.

- abstract: There are no S6+ cations in sulphates. Oxidation state and charge is not the same!

Corrected

- introduction: "thermal destruction method" is an unusual term. Maybe call it "thermal decomposition" or "combustion"?

Replaced by the term "Thermal decomposition"

- results and discussion: For the syntheses the formation of either S or SO2 was proposed. Please explain, how this was proven.

Unfortunately, we have no opportunity to conduct thermogravimetric analysis in a restorative atmosphere. The equations are based on the analysis of a large amount of experimental data. Solid reaction products are described based on x-ray phase analysis data. By-products are determined based on experimental observations in the synthesis process. Elemental sulfur and water are released at the very first stages of synthesis and condense on the cold walls of the reactor (Fig. 1). That allows you to uniquely identify them.

- space group symbols: Please write space group symbols according to the common style as seen e. g. in the International Tables for Crystallography thoughout text and tables.

Corrected

- figure 1 caption: The figure does not only show difference, but also observed and calculated patterns.

Corrected

Reviewer #3: In this manuscript the authors present a method for the production of luminescent materials on the basis of rare earth oxysulfides. The results presented are interesting and the conclusions are adequately supported by the results and discussions presented. After the revisions made by the authors the manuscript is almost ready for publication, I suggest some minor corrections:

The authors should standardize throughout the text "rare-earth" or "rare earth"
Page 1, line 46 the authors should be put a space between "used" and "in"
Page 2, lines 16 and 17 the authors should be replaced "mol/l" to "mol/L"
Page 2, lines 21,22,23 and 27 the authors should be replaced "ml" to "mL"

All comments are taken into account and adjustments are made to the article.

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Synthesis and Optical Properties RE_2O_2S:Ln (RE = La, Y; Ln = Ce, Eu, Dy, Er)

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Abstract

The phase formation sequence was studied in the preparation of solid solutions of RE_2O_2S: Ln' (RE = La, Y; Ln' = Ce, Eu, Dy, Er) by the reduction of the match co-precipitated sulfates, followed by sulfidization of the reduction products. For uniform distribution of cations in the matrix, a method of chemical homogenization was used, consisting in the preparation of an aqueous solution containing all the necessary cations and their subsequent precipitation in the form of sulfates. The use of sulfates as precursors facilitates the process of obtaining solid solutions of oxysulfides, since sulfates already contain SO_4^{2-}-ions. The phase and morphological certification of the obtained solid solutions was carried out. The study of steady state luminescent properties demonstrated characteristic bands which are assigned to *4f-4f* and *5d-4f* transition. The obtained results showed the possibility of applying the method to synthesize optical ceramics based on solid solutionsRE_2O_2S: Ln (RE = La, Y; Ln = Ce, Eu, Dy, Er).

Keywords: oxysulfides, sulfates, rare-earth, reduction, optical ceramic, luminescence

1. Introduction

Oxygen-containing compounds of rare-earth elements have long attracted the attention of researchers due to their effective luminescent properties, which have found application in many optical systems [1-6].Despite the fact that the luminescence is mainly determined by the nature of the substituting ion, the host matrix into which this ion is embedded influences on the emission lines intensity through its crystal field [7-10]. Lanthanide ions can emit light in the near UV, visible and infrared regions of the spectrum. Each ion has a characteristic absorption and emission spectrum. Ln^{3+} radiationis characterized by high color purity; therefore, materials activated by lanthanides are attractive for creating LEDs, fluorescent lamps, plasma displays, and active media for solid-state lasers [11-13].

Activated materials based on oxisulfides of rare-earth elements are widely used in various fields [14-22]. However, in recent years, only laborious, poorly reproducible methods of producing nanoparticles are

described in the literature for these objects. At the same time, the need for relatively simple synthesis methods allowing large batches of optical ceramics does not decrease[23–27].

Thermal decomposition methods are convenient for obtaining materials with different properties [28-32]. In the preparation of oxysulfides, particular attention is drawn to methods for reducing sulfur containing compounds to higher oxidation degrees. Compared with solid-phase methods, the recovery method differs in manufacturability, reproducibility, and the ability to produce several tens of grams of a product at once [33–39].

Thus, the aim of the work is to study the chemistry of reactions in the sequential processing of co-precipitated sulfates of rare earth elements in the atmosphere of H_2, H_2S to obtain activated oxysulfides and investigate the morphology of the obtained reaction products and their luminescent properties.

2. Materials and methods

2a. Preparative Methods

Powders of co-precipitated sulfates were obtained by precipitation from a nitrate solution with concentrated sulfuric acid. For the synthesis, high purity reagents were used: Ln_2O_3 (≥99.99%, ultrapure, TDM-96 Ltd. Russia). Concentrated nitric acid solution ($C(HNO_3)$ = 14.6 mol/L, ultrapure, Vekton Ltd., Russia), concentrated sulfuric acid solution ($C(H_2SO_4)$ = 17.9 mol/L, ultrapure, Vekton Ltd., Russia). Weighing was carried out on an analytical balance with an accuracy of 0.1 mg. Before weighing, the oxides were calcined in a muffle furnace at a temperature of 1000°C for 12 hours to remove sorbed gases and products of their interaction with oxides ($Ln(OH)_3$, $Ln_2(CO_3)_2$). Acid solutions were measured using glass measuring cylinders with an accuracy of 0.1 ml.

The calculated weighed amount of oxides with a total weight of 5.0 g was placed in a 100 mL glass round-bottom flask, then 7.0 mL of a concentrated solution of nitric acid was poured in small portions. The reaction mixture was heated on a mantle until the oxides were completely dissolved. As a result, a nitrate solution was obtained with evenly distributed cations:

$$0.99RE_2O_3 + 0.01Ln_2O_3 + 6HNO_3 \rightarrow 2(RE_{0.99}Ln_{0.01})(NO_3)_3 + 3H_2O \quad (1)$$

After cooling the solution, 3.0 ml of concentrated sulfuric acid was poured in small portions to it, avoiding strong heating of the reaction mixture. As a result, a precipitate of co-precipitated sulfates and their crystalline hydrates forms:

$$2(RE_{0.99}Ln_{0.01})(NO_3)_3 + 3H_2SO_4 + nH_2O \rightarrow (RE_{0.99}Ln_{0.1})_2(SO_4)_3 \cdot nH_2O + 6HNO_3 \quad (2)$$

After carrying out the precipitation reaction, the reaction mixture is distilled off to the dry residue. The obtained polycrystalline product is additionally calcined in a tubular furnace at a temperature of 500°C to remove sorbed moisture and acids. Later, the powder is annealed at the same temperature for 7 days, in order to form an acceptable crystallite structure.

This method of chemical homogenization has a number of significant advantages:

- In the process of synthesis, no cations other than Ln^{3+} are added to the reaction mixture, which excludes their replacement and the formation of defects in the crystal structure.
- Sulfates precipitate from a homogeneous nitrate solution, which ensures high stoichiometry and uniform distribution of cations within the crystal lattice.
- Conducting the reaction in an environment of concentrated sulfuric acid allows to form the structure of anhydrous sulfate at the earliest stages.

The reduction of sulfates in a hydrogen atmosphere was carried out on the apparatus shown in Figure S1. High purity hydrogen was obtained by the electrolytic method in the SPECTR - 6M hydrogen generator. The temperature in the furnace was set using a microprocessor controller. The temperature in the furnace was controlled using chromel-alumel thermocouple. A weighed amount of co-precipitated sulfates was placed in a quartz reactor, and for 30 minutes it was purged with hydrogen from a generator at a rate of 6

L/h. After that, the reactor was placed in a heated vertical furnace and kept for the required amount of time. After completion of the process, the rector was removed from the furnace and cooled to room temperature.
Processing of the reduced products in an atmosphere of hydrogen sulfide was carried out on a similar setup (Fig. S2). The difference lies in the fact that before being fed into the reactor, the hydrogen passes through a flask with molten elemental sulfur and heated to 350 °C. As a result, hydrogen sulfide is formed:
$H_2 + S \rightarrow H_2S$ (3)
Consequently, it is not a hydrogen inflows the reactor, but a hydrogen sulfide.

2b. Methods of physical-chemical analysis

X-ray phase analysis (XRD) was performed on a BRUKER D2 PHASER diffractometer with a linear detector LYNXEYE (CuKα radiation, Ni-filter). Rietveld refinement of all six samples was performed by using TOPAS 4.2 [40]. Almost all peaks were indexed.
Electron-microscopic analysis was carried out on electron microscope JEOL JSM-6510LV. X-ray energy-dispersive analyzer was used to register X-rays at element spectrum plotting in selected sample surface areas. The inaccuracy in element content determination was equal to ±0.2%.
All measurements of the luminescent properties were carried out on a research-grade spectrofluorimeter. Horiba JobinYvon Fluorolog-3 equipped with double monochromators for excitation and emission channels and 450 W xenon lamp as an excitation source.

3. Result and Discussion

3a. Synthetic experiment

Detailed consideration of the chemical transformations taking place during the transformation of co-precipitated sulfates into the corresponding solid solutions of oxysulfides was made on the basis of two model systems $La_2(SO_4)_3$:Dy^{3+} and $Y_2(SO_4)_3$:Er^{3+}. The results obtained were used to synthesize all other solid solutions, which are reported in this work.
The carrying out of the co-precipitation of sulfates and the subsequent annealing led to the formation of structures of solid solutions of sulfates in which the doping ion is fullyincorporated into the crystal lattice of the matrix and occupies the crystallographic positions of the host cation. According to X-ray diffraction data, all samples of co-precipitated sulfates are single-phase (Fig. 1a, b). There is a slight shift in the unit cell parameters caused by the difference in the radii of the cations of the matrix and the dopant.
The appearance of gaseous reduction products was recorded at a temperature of 570°C. In this connection, sulfate reduction was carried out at t = 600°C. At this temperature, after 60 minutes of the process, the products mainly consist of 4 phases: $(RE_{0.99}Ln_{0.01})_2(SO_4)_3$- $(RE_{0.99}Ln_{0.01})_2O_2SO_4$ - $(RE_{0.99}Ln_{0.01})_2O_2S$ - $(RE_{0.99}Ln_{0.01})_2O_3$ (Fig. 1c, d).There was an incomplete transformation of the initial sulfates into the reaction products. The following chemical equations correspond to the formation of the corresponding reduction products:
$(RE_{0.99}Ln_{0.01})_2(SO_4)_3 + 6H_2 \rightarrow (RE_{0.99}Ln_{0.01})_2O_2SO_4 + 2S + 6H_2O$ (4)
$(RE_{0.99}Ln_{0.01})_2(SO_4)_3 + 10H_2 \rightarrow (RE_{0.99}Ln_{0.01})_2O_2S + 2S + 10H_2O$ (5)
$(RE_{0.99}Ln_{0.01})_2(SO_4)_3 + 5H_2 \rightarrow (RE_{0.99}Ln_{0.01})_2O_3 + 2SO_2 + 5H_2O$ (6)
$(RE_{0.99}Ln_{0.01})_2O_2SO_4 + 4H_2 \rightarrow (RE_{0.99}Ln_{0.01})_2O_2S + 4H_2O$ (7)
After 10 hours of carrying out the process at a given temperature, polycrystalline products consist of two phases: $(RE_{0.99}Ln_{0.01})_2O_2S$, $(RE_{0.99}Ln_{0.01})_2O_3$ (Fig. 1 e, f). The absence of compounds containing sulfur in the highest degree of oxidation of in the synthesis products indicates on a complete redox transformation. In all samples, the phase output $(RE_{0.99}Ln_{0.01})_2O_2S$ was not lower than 80%. Thus, the stage of sulfate reduction in a hydrogen atmosphere allows the formation of two-phase polycrystalline intermediates with a predominant content of the oxysulfide phase, which should greatly facilitate the sulfidation procedure.

The interaction of two-phase intermediates with hydrogen sulfide at a temperature of 800 ° C for 5 hours leads to the formation of single-phase powders $(RE_{0.99}Ln_{0.01})_2O_2S$ (Fig. 1 g, h). The transformation corresponds to the transformation of the oxide phase to the oxysulfide under the action of a mild sulfiding agent H_2S:

$(RE_{0.99}Ln_{0.01})_2O_3 + H_2S \rightarrow (RE_{0.99}Ln_{0.01})_2O_2S + H_2O$ (8)

Sulfate powders, formed predominantly by loose agglomerates (Fig. 2, a) with sizes up to 10 μm. The resulting oxysulfide powders have a more distinct cut shape, a denser structure and a uniform size distribution (Fig. 2, a). The morphological transformation is evidently due to the elevated temperatures and the diffusion character of the reduction and sulfidation processes.

The enlargement obviously occurs as a result of the desire of the system to lower its energy. What corresponds to a decrease in the surface area of the polycrystalline samples. Particle cutting appears as a result of high rates of chemical reactions and rapid mass transfer. The indicated tendency to particle aggregation is often observed during similar processes [27, 32-35].

3b. Structural and spectroscopic properties

Crystal structure of both hosts belongs to *P-3m 1* space group of trigonal symmetry class. Y and La occupy a single inequivalent site. In both oxysulfides the local environment of them is a distorted polyhedron with seven vertices, four of them being oxygen ions and three being sulfur ions. Layered structure of oxysulfides implies that sulfur and oxygen are positioned in opposite hemispheres of the local environment of either Y or La. Rare-earth doping ions are expected to occupy Y and La sites, and their local environment is determined by the structure host, with the local symmetry C_{3v}. Therefore, absence of inversion symmetry must be pronounced in optical spectra of doping ions. Variation of luminescence properties of doping RE ions in one host with respect to another is usually ascribed to the change of the extent of inversion symmetry violation. Examining the geometry of local environment of RE ion in Y_2O_2S and La_2O_2S (Fig. 3) we observe that both environments are geometrically identical.

The excitation and emission spectra of RE_2O_2S (RE = Y, La) activated by 1% of Dy^{3+} ions are shown in Fig. 4. The observed spectra exhibit characteristic intra-configurational 4f-4f transitions. Excitation spectrum of $Y_2O_2S:Dy^{3+}$ monitored at 579 nm ($^4F_{9/2}$–$^6H_{13/2}$), displays following transitions: $^6H_{15/2}$–$^4P_{7/2}$ (355 nm), $^6H_{15/2}$–$^4P_{5/2}$ (369 nm), $^6H_{15/2}$–$^4I_{13/2}$(388 nm), $^6H_{15/2}$–$^4G_{11/2}$ (427 nm), $^6H_{15/2}$–$^4I_{15/2}$ (451 nm) and $^6H_{15/2}$–$^4F_{9/2}$ (479 nm). The emission spectrum $Y_2O_2S:Dy^{3+}$ sample is dominated by green-yellow band (579 nm) corresponding to the hypersensitive $^4F_{9/2}$–$^6H_{13/2}$ transition. Other observed lines are attributed to the $^4I_{15/2}$–$^6H_{15/2}$ (457 nm), $^4F_{9/2}$–$^6H_{15/2}$ (487 nm), and$^4F_{9/2}$–$^6H_{11/2}$ (670 nm) transitions. It is well-known that $^4F_{9/2}$–$^6H_{13/2}$ is the forced electric dipole transition, which is hypersensitive and its intensity can vary by orders of magnitude depending on the local site symmetry, whereas $^4F_{9/2}$–$^6H_{15/2}$ transition intensity is insignificantly affected by the environment [40-42]. The excitation and emission spectra of $La_2O_2S:Dy^{3+}$ are similar to the $Y_2O_2S:Dy^{3+}$ ones. Small blue shift of bands and redistribution between them were observed. So, the most prominent transitions in excitation and emission spectra of $La_2O_2S:Dy^{3+}$are centered at 353 and 576 nm, respectively. Observed luminescence spectra of Dy ion are consistent with the concept that they occupy Y(La) sites with the local symmetry C_{3v}.

To compare the crystal structure and crystal field of $Y_2O_2S:Dy^{3+}$ and $La_2O_2S:Dy^{3+}$ powders, we calculated ratio (R_{Dy}) between $^4F_{9/2}$–$^6H_{13/2}$ and $^4F_{9/2}$–$^6H_{15/2}$ intensities. This parameter is similar to the well-known asymmetry ratio for Eu^{3+} ions [43, 44]. R_{Dy} value give information about the local surrounding and environmental changes near the Dy^{3+} ions. The higher the calculated parameter is, the more apart from a centrosymmetric geometryluminescent center is located. It is well-known that if Dy^{3+} is located at low symmetry without the inversion symmetry, the yellow emission is the most intense of all the transitions, as is the case with our synthesized nanocrystalline phosphors [45]. Experimental R_{Dy} values for $Y_2O_2S:Dy^{3+}$

and $La_2O_2S:Dy^{3+}$ samples are 3.13 and 4.65. In view of geometrical identity of local environments, this difference must be ascribed to the interplay between ionic radii of Y and La and the unit cell parameters, i.e. closer ligands at the same degree of inversion symmetry violation.

The steady state luminescence spectra of RE_2O_2S (RE = Y, La) powders doped with 1% of Er^{3+} ions are presented in Fig. 5. The excitation spectrum of $Y_2O_2S:Er^{3+}$was monitored at 549 nm($^4F_{9/2}$–$^4I_{15/2}$) within spectral range of 330–530 nm. It consists of $^4I_{15/2}$–$^4G_{7/2}$ (360 nm), $^4I_{15/2}$–$^4G_{9/2}$ (367 nm), $^4I_{15/2}$–$^4G_{11/2}$ (379 nm), $^4I_{15/2}$–$^2H_{9/2}$ (408 nm), $^4I_{15/2}$–$^4F_{3/2}$ (446 nm),$^4I_{15/2}$–$^4F_{5/2}$ (453 nm),$^4I_{15/2}$–$^4F_{7/2}$ (491 nm), and $^4I_{15/2}$–$^2H_{11/2}$ (522 nm).The emission spectrum includes narrow bands, which are assigned to the following transitions: $^2H_{9/2}$–$^4I_{15/2}$ (409 nm), $^2H_{11/2}$–$^4I_{15/2}$ (524 nm), $^4S_{3/2}$–$^4I_{15/2}$ (549 nm), and $^4F_{9/2}$–$^4I_{15/2}$ (670 nm).The spectral line positions of $La_2O_2S:Er^{3+}$ spectra are the same. Change of host leads to the intensity redistribution, which is most pronounced for $^4I_{15/2}$–$^4G_{11/2}$ transition in the excitation spectrum.

Fig. 6 displays excitation and emission spectra of RE_2O_2S (RE = Y, La) activated by 1% of Eu^{3+} ions. The excitation spectrum of $Y_2O_2S:Eu^{3+}$ monitored at 545 nm ($^5D_1$–$^7F_1$) consists of following transitions:$^7F_0$–$^5D_4$ (353 nm),$^7F_0$–$^5L_7$(378 nm) and $^7F_2$–$^5D_2$(488 nm). The emission spectrum shows narrow bands originating from $^5D_1$and $^5D_0$ excited levels. Surprisingly, that the emission spectrum of $Y_2O_2S:Eu^{3+}$ is dominated by transition$^5D_1$–$^7F_1$(545 nm), whereas the most prominent luminescence bands are usually attributed to the $^5D_0$–^7F_J transition [46-48]. Such behavior was previously reported for $La_2O_2S:Eu^{3+}$ bulk phosphors [49, 50].

Dominance of $^5D_1$ emission can be explained by small phonon energy in regarded host, because significant amount of ionsrelax to $^5D_1$ level after the UV excitation , and they radiatively decay to the ground state before nonradiative decay to $^5D_0$ metastable level. We also observed $^5D_2$–$^7F_2$(490 nm), $^5D_1$–$^7F_3$ (587 nm), $^5D_0$–$^7F_1$ (592 nm), $^5D_0$–$^7F_2$ (621 nm) and $^5D_0$–$^7F_3$ (670 nm) transitions.

The excitation and emission spectra of $La_2O_2S:Eu^{3+}$ display situation, which is more usual for Eu^{3+}-doped compounds. The excitation spectrum of $Y_2O_2S:Eu^{3+}$ monitored at 623 nm ($^5D_0$–$^7F_2$) consists of broad intense band corresponding to charge transfer S^{2-}–Eu^{3+} (338 nm) and low-intense line assigned to the typical intra-configurational transitionsof the Eu^{3+} ion: $^7F_0$–$^5L_6$ (394 nm), $^7F_0$–$^5D_2$ (466 nm), $^7F_0$–$^5D_1$ (536 nm) and $^7F_1$–$^5D_0$ (593 nm). The emission spectrum is dominated by the forced electric dipole transition $^5D_0$–$^7F_2$ with maximum at 623 nm. Other observed lines are attributed to the $^5D_1$–$^7F_1$ (538 nm), $^5D_1$–$^7F_2$ (555 nm), $^5D_1$–$^7F_3$ (586 nm), $^5D_0$–$^7F_1$ (594 nm), $^5D_0$–$^7F_3$ (670 nm) and $^5D_0$–$^7F_4$ (704 nm).

Due to the unique luminescence properties of Eu^{3+} ions, it is quite easy to analyze the luminescent center local surrounding and its symmetry using only emission spectrum. The asymmetry ratio (R_{Eu}) gives information about local changes around the Eu^{3+} ions. It is defined as intensity ratio of forced electric dipole $^5D_0$–$^7F_2$ and magnetic dipole $^5D_0$–$^7F_1$ transitions. The higher the asymmetry parameter R_{Eu} is, the more apart from a centrosymmetric geometry luminescent center is located. The calculated R_{Eu} values of $Y_2O_2S:Eu^{3+}$ and $La_2O_2S:Eu^{3+}$ samples are 0.58 and 3.16, respectively. It is worth noting that the calculated R_{Eu} values of $Y_2O_2S:Eu^{3+}$ and $La_2O_2S:Eu^{3+}$ samples significantly differ, which indicate big difference in local surrounding of Eu^{3+} ions in these hosts.

The steady state luminescence spectra of RE_2O_2S (RE = Y, La) powders doped with 1% of Ce^{3+} ions are shown in Fig. 7. The excitation spectrum of $Y_2O_2S:Ce^{3+}$ sample displays two broad bands centered at 265 and 407 nm (λ_{em} = 545 nm). These bands correspond to direct excitation of the Ce^{3+} ions via transitions to the components of Ce^{3+} 5d configuration. The emission spectrum also consists of two lines attributed to allowed 5d–4f transition of Ce^{3+} ion. Generally, emission lines attributed to allowed 5d–4f transition in Ce^{3+}-doped materials are quite broad [51, 52]. Sometimes they are split into two components separated by approximately 2000 cm^{-1} due to the spin-orbit splitting of the $4f^1$ ground state into two components $^2F_{5/2}$ and $^2F_{7/2}$. The bands observed in $Y_2O_2S:Ce^{3+}$ and $La_2O_2S:Ce^{3+}$ exhibit splitting by 5000 cm^{-1} and cannot be due to splitting of the ground stated mentioned above. Therefore, two bands ib Ce^{3+} luminescence must be ascribed to the electron transitions from the lowest and second 5d levels to the ground state of Ce^{3+}[53].

Change of host to La_2O_2S does not affect spectroscopic properties of Ce^{3+}-doped material. The line positions are almost the same for both excitation and emission spectra.

4. Conclusions

In summary, a method for the production of luminescent materials on the basis of rare-earth oxysulfides was developed. The advantage of the method consists in the precipitation of a sulfur-containing precursor from a homogeneous nitrate solution and subsequent transformation in a reducing and sulfidating atmosphere. The use of chemical homogenization made it possible to achieve an excellent uniform distribution of cations in the structure. The use of sulfates as precursors, in view of the presence of sulfur in the structure, greatly simplifies the process of obtaining solid solutions of oxysulfides. All synthesized samples have single phase without any impurities. The excitation and emission spectra of RE_2O_2S:Ln (RE = Y, La; Ln = Dy, Er, Eu) consist of characteristic bands corresponding to the 4f-4f intra configurational transitions. The study of Dy^{3+} and Eu^{3+}-doped powders revealed that Y_2O_2S host possesses higher local symmetry than La_2O_2S one. The excitation and emission spectra of $RE_2O_2S:Ce^{3+}$ (RE = Y, La) phosphor displayed allowed 5d–4f transition.

Acknowledgement

The authors would like to thank the staff of the Engineering Center of the Tyumen State University for their help in carrying out physical and chemical tests. We also thank Andrey Bobylev for conducting electron microscopy.

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Captions

Fig.1. Experimental, calculated, and difference Rietveld plot of: a,b) $(RE_{0.99}Ln_{0.01})_2(SO_4)_3$; c,d) $(RE_{0.99}Ln_{0.01})_2(SO_4)_3$-$(RE_{0.99}Ln_{0.01})_2O_2SO_4$-$(RE_{0.99}Ln_{0.01})_2O_2S$-$(RE_{0.99}Ln_{0.01})_2O_3$; e,f) $(RE_{0.99}Ln_{0.01})_2O_2S$ - $(RE_{0.99}Ln_{0.01})_2O_3$; g,h) $(RE_{0.99}Ln_{0.01})_2O_2S$.

Fig. 2. SEM image of a) $(La_{0.99}Dy_{0.01})_2(SO_4)_3$; b) $(La_{0.99}Dy_{0.01})_2O_2S$

Fig. 3. Coordination polyhedron structure La_2O_2S

Fig. 4. Excitation and emission spectra of $Y_2O_2S:Dy^{3+}$and $La_2O_2S:Dy^{3+}$phosphors

Fig. 5. Excitation and emission spectra of $Y_2O_2S:Er^{3+}$ and $La_2O_2S:Er^{3+}$phosphors

Fig. 6. Excitation and emission spectra of $Y_2O_2S:Eu^{3+}$ and $La_2O_2S:Eu^{3+}$phosphors

Fig. 7. Excitation and emission spectra of $Y_2O_2S:Ce^{3+}$ and $La_2O_2S:Ce^{3+}$phosphors

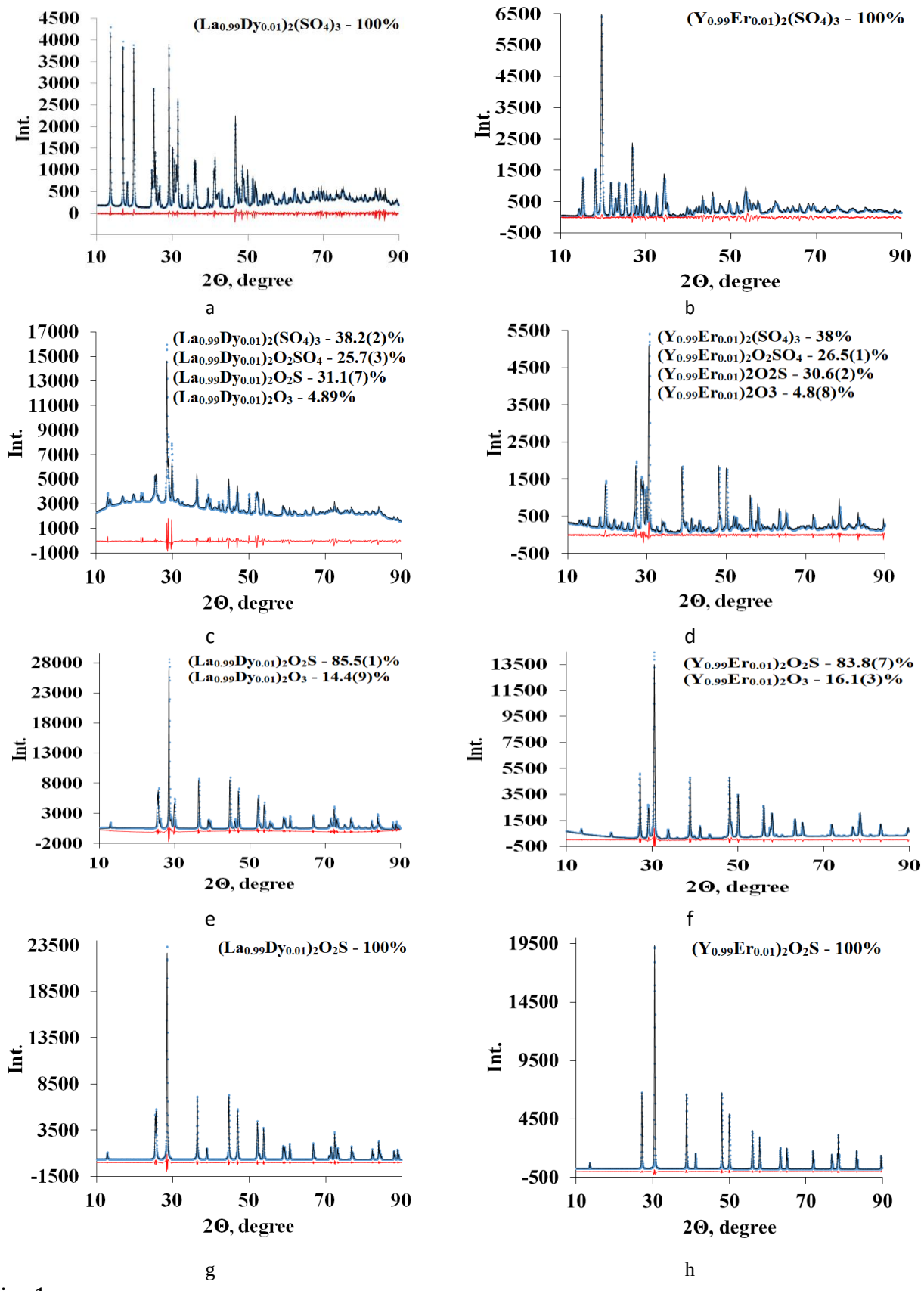
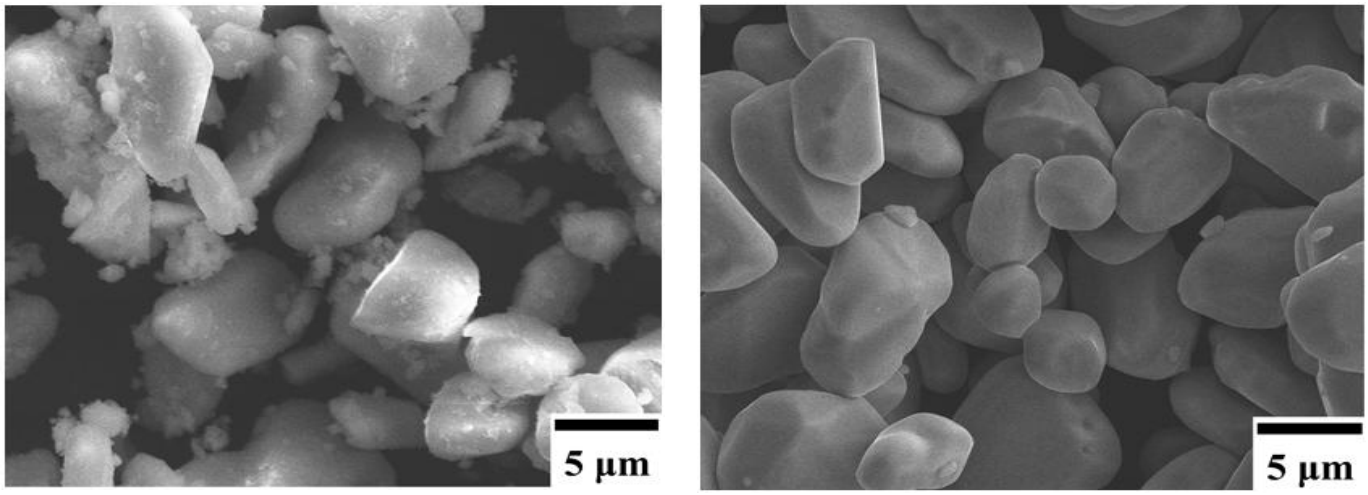

(La0.99Dy0.01)2(SO4)3 - 100%
Int.
2Θ, degree
a
(Y0.99Er0.01)2(SO4)3 - 100%
Int.
2Θ, degree
b
(La0.99Dy0.01)2(SO4)3 - 38.2(2)%
(La0.99Dy0.01)2O2SO4 - 25.7(3)%
(La0.99Dy0.01)2O2S - 31.1(7)%
(La0.99Dy0.01)2O3 - 4.89%
Int.
2Θ, degree
c
(Y0.99Er0.01)2(SO4)3 - 38%
(Y0.99Er0.01)2O2SO4 - 26.5(1)%
(Y0.99Er0.01)2O2S - 30.6(2)%
(Y0.99Er0.01)2O3 - 4.8(8)%
Int.
2Θ, degree
d
(La0.99Dy0.01)2O2S - 85.5(1)%
(La0.99Dy0.01)2O3 - 14.4(9)%
Int.
2Θ, degree
e
(Y0.99Er0.01)2O2S - 83.8(7)%
(Y0.99Er0.01)2O3 - 16.1(3)%
Int.
2Θ, degree
f
(La0.99Dy0.01)2O2S - 100%
Int.
2Θ, degree
g
(Y0.99Er0.01)2O2S - 100%
Int.
2Θ, degree
h

Fig. 1



a) b)

Fig. 2

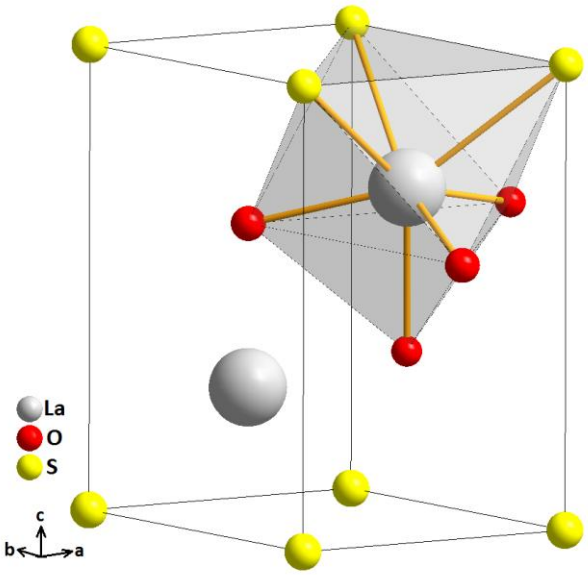


Fig. 3.

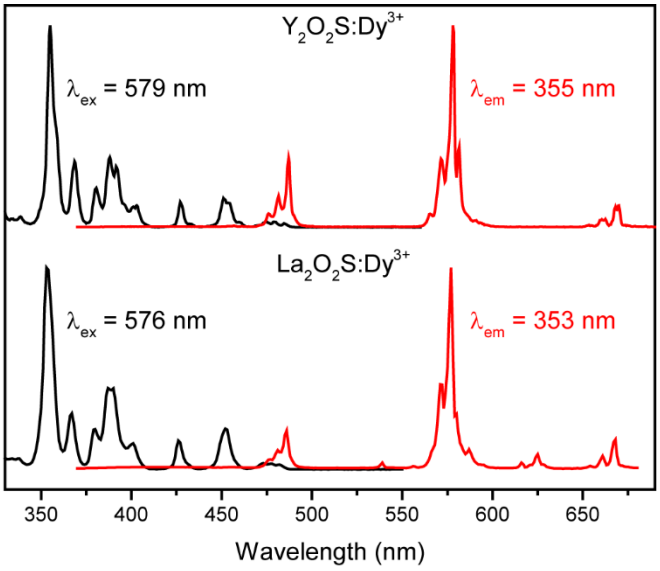


Fig. 4

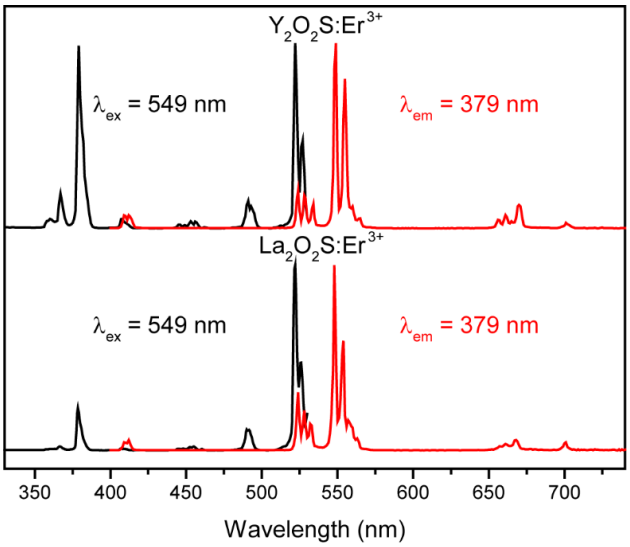


Fig. 5

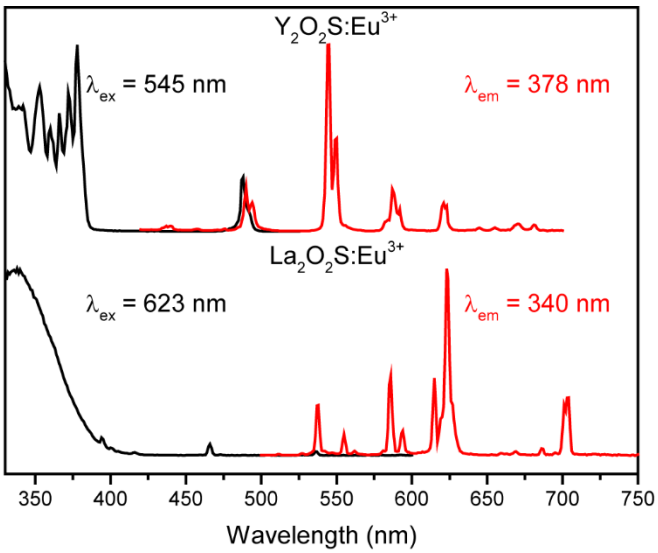


Fig. 6

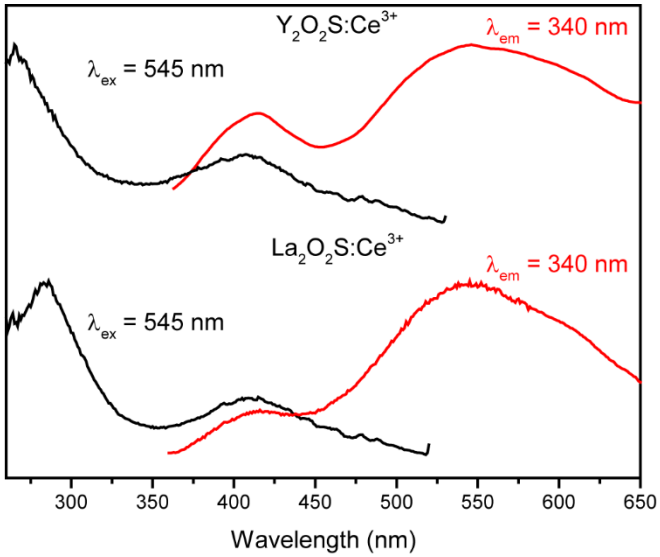

$Y_2O_2S:Ce^{3+}$
λ_{em} = 340 nm
λ_{ex} = 545 nm
$La_2O_2S:Ce^{3+}$
λ_{em} = 340 nm
λ_{ex} = 545 nm
300
350
400
450
500
550
600
650
Wavelength (nm)

Fig. 7

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Synthesis and Optical Properties RE_2O_2S:Ln (RE = La, Y; Ln = Ce, Eu, Dy, Er)

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Abstract

The phase formation sequence was studied in the preparation of solid solutions of RE_2O_2S: Ln' (RE = La, Y; Ln' = Ce, Eu, Dy, Er) by the reduction of the match co-precipitated sulfates, followed by sulfidization of the reduction products. For uniform distribution of cations in the matrix, a method of chemical homogenization was used, consisting in the preparation of an aqueous solution containing all the necessary cations and their subsequent precipitation in the form of sulfates. The use of sulfates as precursors facilitates the process of obtaining solid solutions of oxysulfides, since sulfates already contain SO_4^{2-}-ions. The phase and morphological certification of the obtained solid solutions was carried out. The study of steady state luminescent properties demonstrated characteristic bands which are assigned to *4f-4f* and *5d-4f* transition. The obtained results showed the possibility of applying the method to synthesize optical ceramics based on solid solutionsRE_2O_2S: Ln (RE = La, Y; Ln = Ce, Eu, Dy, Er).

Keywords: oxysulfides, sulfates, rare-earth, reduction, optical ceramic, luminescence

1. Introduction

Oxygen-containing compounds of rare-earth elements have long attracted the attention of researchers due to their effective luminescent properties, which have found application in many optical systems [1-6].Despite the fact that the luminescence is mainly determined by the nature of the substituting ion, the host matrix into which this ion is embedded influences on the emission lines intensity through its crystal field [7-10]. Lanthanide ions can emit light in the near UV, visible and infrared regions of the spectrum. Each ion has a characteristic absorption and emission spectrum. Ln^{3+} radiationis characterized by high color purity; therefore, materials activated by lanthanides are attractive for creating LEDs, fluorescent lamps, plasma displays, and active media for solid-state lasers [11-13].

Activated materials based on oxisulfides of rare-earth elements are widely used in various fields [14-22]. However, in recent years, only laborious, poorly reproducible methods of producing nanoparticles are

described in the literature for these objects. At the same time, the need for relatively simple synthesis methods allowing large batches of optical ceramics does not decrease[23–27].

Thermal decomposition methods are convenient for obtaining materials with different properties [28-32]. In the preparation of oxysulfides, particular attention is drawn to methods for reducing sulfur containing compounds to higher oxidation degrees. Compared with solid-phase methods, the recovery method differs in manufacturability, reproducibility, and the ability to produce several tens of grams of a product at once [33–39].

Thus, the aim of the work is to study the chemistry of reactions in the sequential processing of co-precipitated sulfates of rare earth elements in the atmosphere of H_2, H_2S to obtain activated oxysulfides and investigate the morphology of the obtained reaction products and their luminescent properties.

2. Materials and methods

2a. Preparative Methods

Powders of co-precipitated sulfates were obtained by precipitation from a nitrate solution with concentrated sulfuric acid. For the synthesis, high purity reagents were used: Ln_2O_3 (≥99.99%, ultrapure, TDM-96 Ltd. Russia). Concentrated nitric acid solution (C(HNO_3) = 14.6 mol/L, ultrapure, Vekton Ltd., Russia), concentrated sulfuric acid solution (C (H_2SO_4) = 17.9 mol/L, ultrapure, Vekton Ltd., Russia). Weighing was carried out on an analytical balance with an accuracy of 0.1 mg. Before weighing, the oxides were calcined in a muffle furnace at a temperature of 1000°C for 12 hours to remove sorbed gases and products of their interaction with oxides ($Ln(OH)_3$, $Ln_2(CO_3)_2$). Acid solutions were measured using glass measuring cylinders with an accuracy of 0.1 ml.

The calculated weighed amount of oxides with a total weight of 5.0 g was placed in a 100 mL glass round-bottom flask, then 7.0 mL of a concentrated solution of nitric acid was poured in small portions. The reaction mixture was heated on a mantle until the oxides were completely dissolved. As a result, a nitrate solution was obtained with evenly distributed cations:

$$0.99RE_2O_3 + 0.01Ln_2O_3 + 6HNO_3 \rightarrow 2(RE_{0.99}Ln_{0.01})(NO_3)_3 + 3H_2O \quad (1)$$

After cooling the solution, 3.0 ml of concentrated sulfuric acid was poured in small portions to it, avoiding strong heating of the reaction mixture. As a result, a precipitate of co-precipitated sulfates and their crystalline hydrates forms:

$$2(RE_{0.99}Ln_{0.01})(NO_3)_3 + 3H_2SO_4 + nH_2O \rightarrow (RE_{0.99}Ln_{0.1})_2(SO_4)_3 \cdot nH_2O + 6HNO_3 \quad (2)$$

After carrying out the precipitation reaction, the reaction mixture is distilled off to the dry residue. The obtained polycrystalline product is additionally calcined in a tubular furnace at a temperature of 500°C to remove sorbed moisture and acids. Later, the powder is annealed at the same temperature for 7 days, in order to form an acceptable crystallite structure.

This method of chemical homogenization has a number of significant advantages:

- In the process of synthesis, no cations other than Ln^{3+} are added to the reaction mixture, which excludes their replacement and the formation of defects in the crystal structure.
- Sulfates precipitate from a homogeneous nitrate solution, which ensures high stoichiometry and uniform distribution of cations within the crystal lattice.
- Conducting the reaction in an environment of concentrated sulfuric acid allows to form the structure of anhydrous sulfate at the earliest stages.

The reduction of sulfates in a hydrogen atmosphere was carried out on the apparatus shown in Figure S1. High purity hydrogen was obtained by the electrolytic method in the SPECTR - 6M hydrogen generator. The temperature in the furnace was set using a microprocessor controller. The temperature in the furnace was controlled using chromel-alumel thermocouple. A weighed amount of co-precipitated sulfates was placed in a quartz reactor, and for 30 minutes it was purged with hydrogen from a generator at a rate of 6

L/h. After that, the reactor was placed in a heated vertical furnace and kept for the required amount of time. After completion of the process, the rector was removed from the furnace and cooled to room temperature.
Processing of the reduced products in an atmosphere of hydrogen sulfide was carried out on a similar setup (Fig. S2). The difference lies in the fact that before being fed into the reactor, the hydrogen passes through a flask with molten elemental sulfur and heated to 350 °C. As a result, hydrogen sulfide is formed:
$H_2 + S \rightarrow H_2S$ (3)
Consequently, it is not a hydrogen inflows the reactor, but a hydrogen sulfide.

2b. Methods of physical-chemical analysis

X-ray phase analysis (XRD) was performed on a BRUKER D2 PHASER diffractometer with a linear detector LYNXEYE (CuKα radiation, Ni-filter). Rietveld refinement of all six samples was performed by using TOPAS 4.2 [40]. Almost all peaks were indexed.
Electron-microscopic analysis was carried out on electron microscope JEOL JSM-6510LV. X-ray energy-dispersive analyzer was used to register X-rays at element spectrum plotting in selected sample surface areas. The inaccuracy in element content determination was equal to ±0.2%.
All measurements of the luminescent properties were carried out on a research-grade spectrofluorimeter. Horiba JobinYvon Fluorolog-3 equipped with double monochromators for excitation and emission channels and 450 W xenon lamp as an excitation source.

3. Result and Discussion

3a. Synthetic experiment

Detailed consideration of the chemical transformations taking place during the transformation of co-precipitated sulfates into the corresponding solid solutions of oxysulfides was made on the basis of two model systems $La_2(SO_4)_3$:Dy^{3+} and $Y_2(SO_4)_3$:Er^{3+}. The results obtained were used to synthesize all other solid solutions, which are reported in this work.
The carrying out of the co-precipitation of sulfates and the subsequent annealing led to the formation of structures of solid solutions of sulfates in which the doping ion is fullyincorporated into the crystal lattice of the matrix and occupies the crystallographic positions of the host cation. According to X-ray diffraction data, all samples of co-precipitated sulfates are single-phase (Fig. 1a, b). There is a slight shift in the unit cell parameters caused by the difference in the radii of the cations of the matrix and the dopant.
The appearance of gaseous reduction products was recorded at a temperature of 570°C. In this connection, sulfate reduction was carried out at t = 600°C. At this temperature, after 60 minutes of the process, the products mainly consist of 4 phases: $(RE_{0.99}Ln_{0.01})_2(SO_4)_3$- $(RE_{0.99}Ln_{0.01})_2O_2SO_4$ - $(RE_{0.99}Ln_{0.01})_2O_2S$ - $(RE_{0.99}Ln_{0.01})_2O_3$ (Fig. 1c, d).There was an incomplete transformation of the initial sulfates into the reaction products. The following chemical equations correspond to the formation of the corresponding reduction products:
$(RE_{0.99}Ln_{0.01})_2(SO_4)_3 + 6H_2 \rightarrow (RE_{0.99}Ln_{0.01})_2O_2SO_4 + 2S + 6H_2O$ (4)
$(RE_{0.99}Ln_{0.01})_2(SO_4)_3 + 10H_2 \rightarrow (RE_{0.99}Ln_{0.01})_2O_2S + 2S + 10H_2O$ (5)
$(RE_{0.99}Ln_{0.01})_2(SO_4)_3 + 5H_2 \rightarrow (RE_{0.99}Ln_{0.01})_2O_3 + 2SO_2 + 5H_2O$ (6)
$(RE_{0.99}Ln_{0.01})_2O_2SO_4 + 4H_2 \rightarrow (RE_{0.99}Ln_{0.01})_2O_2S + 4H_2O$ (7)
After 10 hours of carrying out the process at a given temperature, polycrystalline products consist of two phases: $(RE_{0.99}Ln_{0.01})_2O_2S$, $(RE_{0.99}Ln_{0.01})_2O_3$ (Fig. 1 e, f). The absence of compounds containing sulfur in the highest degree of oxidation of in the synthesis products indicates on a complete redox transformation. In all samples, the phase output $(RE_{0.99}Ln_{0.01})_2O_2S$ was not lower than 80%. Thus, the stage of sulfate reduction in a hydrogen atmosphere allows the formation of two-phase polycrystalline intermediates with a predominant content of the oxysulfide phase, which should greatly facilitate the sulfidation procedure.

The interaction of two-phase intermediates with hydrogen sulfide at a temperature of 800 ° C for 5 hours leads to the formation of single-phase powders $(RE_{0.99}Ln_{0.01})_2O_2S$ (Fig. 1 g, h). The transformation corresponds to the transformation of the oxide phase to the oxysulfide under the action of a mild sulfiding agent H_2S:

$(RE_{0.99}Ln_{0.01})_2O_3 + H_2S \rightarrow (RE_{0.99}Ln_{0.01})_2O_2S + H_2O$ (8)

Sulfate powders, formed predominantly by loose agglomerates (Fig. 2, a) with sizes up to 10 μm. The resulting oxysulfide powders have a more distinct cut shape, a denser structure and a uniform size distribution (Fig. 2, a). The morphological transformation is evidently due to the elevated temperatures and the diffusion character of the reduction and sulfidation processes.

The enlargement obviously occurs as a result of the desire of the system to lower its energy. What corresponds to a decrease in the surface area of the polycrystalline samples. Particle cutting appears as a result of high rates of chemical reactions and rapid mass transfer. The indicated tendency to particle aggregation is often observed during similar processes [27, 32-35].

3b. Structural and spectroscopic properties

Crystal structure of both hosts belongs to *P-3m 1* space group of trigonal symmetry class. Y and La occupy a single inequivalent site. In both oxysulfides the local environment of them is a distorted polyhedron with seven vertices, four of them being oxygen ions and three being sulfur ions. Layered structure of oxysulfides implies that sulfur and oxygen are positioned in opposite hemispheres of the local environment of either Y or La. Rare-earth doping ions are expected to occupy Y and La sites, and their local environment is determined by the structure host, with the local symmetry C_{3v}. Therefore, absence of inversion symmetry must be pronounced in optical spectra of doping ions. Variation of luminescence properties of doping RE ions in one host with respect to another is usually ascribed to the change of the extent of inversion symmetry violation. Examining the geometry of local environment of RE ion in Y_2O_2S and La_2O_2S (Fig. 3) we observe that both environments are geometrically identical.

The excitation and emission spectra of RE_2O_2S (RE = Y, La) activated by 1% of Dy^{3+} ions are shown in Fig. 4. The observed spectra exhibit characteristic intra-configurational 4f-4f transitions. Excitation spectrum of $Y_2O_2S:Dy^{3+}$ monitored at 579 nm ($^4F_{9/2}$–$^6H_{13/2}$), displays following transitions: $^6H_{15/2}$–$^4P_{7/2}$ (355 nm), $^6H_{15/2}$–$^4P_{5/2}$ (369 nm), $^6H_{15/2}$–$^4I_{13/2}$(388 nm), $^6H_{15/2}$–$^4G_{11/2}$ (427 nm), $^6H_{15/2}$–$^4I_{15/2}$ (451 nm) and $^6H_{15/2}$–$^4F_{9/2}$ (479 nm). The emission spectrum $Y_2O_2S:Dy^{3+}$ sample is dominated by green-yellow band (579 nm) corresponding to the hypersensitive $^4F_{9/2}$–$^6H_{13/2}$ transition. Other observed lines are attributed to the $^4I_{15/2}$–$^6H_{15/2}$ (457 nm), $^4F_{9/2}$–$^6H_{15/2}$ (487 nm), and$^4F_{9/2}$–$^6H_{11/2}$ (670 nm) transitions. It is well-known that $^4F_{9/2}$–$^6H_{13/2}$ is the forced electric dipole transition, which is hypersensitive and its intensity can vary by orders of magnitude depending on the local site symmetry, whereas $^4F_{9/2}$–$^6H_{15/2}$ transition intensity is insignificantly affected by the environment [40-42]. The excitation and emission spectra of $La_2O_2S:Dy^{3+}$ are similar to the $Y_2O_2S:Dy^{3+}$ ones. Small blue shift of bands and redistribution between them were observed. So, the most prominent transitions in excitation and emission spectra of $La_2O_2S:Dy^{3+}$are centered at 353 and 576 nm, respectively. Observed luminescence spectra of Dy ion are consistent with the concept that they occupy Y(La) sites with the local symmetry C_{3v}.

To compare the crystal structure and crystal field of $Y_2O_2S:Dy^{3+}$ and $La_2O_2S:Dy^{3+}$ powders, we calculated ratio (R_{Dy}) between $^4F_{9/2}$–$^6H_{13/2}$ and $^4F_{9/2}$–$^6H_{15/2}$ intensities. This parameter is similar to the well-known asymmetry ratio for Eu^{3+} ions [43, 44]. R_{Dy} value give information about the local surrounding and environmental changes near the Dy^{3+} ions. The higher the calculated parameter is, the more apart from a centrosymmetric geometryluminescent center is located. It is well-known that if Dy^{3+} is located at low symmetry without the inversion symmetry, the yellow emission is the most intense of all the transitions, as is the case with our synthesized nanocrystalline phosphors [45]. Experimental R_{Dy} values for $Y_2O_2S:Dy^{3+}$

and $La_2O_2S:Dy^{3+}$ samples are 3.13 and 4.65. In view of geometrical identity of local environments, this difference must be ascribed to the interplay between ionic radii of Y and La and the unit cell parameters, i.e. closer ligands at the same degree of inversion symmetry violation.

The steady state luminescence spectra of RE_2O_2S (RE = Y, La) powders doped with 1% of Er^{3+} ions are presented in Fig. 5. The excitation spectrum of $Y_2O_2S:Er^{3+}$was monitored at 549 nm($^4F_{9/2}$–$^4I_{15/2}$) within spectral range of 330–530 nm. It consists of $^4I_{15/2}$–$^4G_{7/2}$ (360 nm), $^4I_{15/2}$–$^4G_{9/2}$ (367 nm), $^4I_{15/2}$–$^4G_{11/2}$ (379 nm), $^4I_{15/2}$–$^2H_{9/2}$ (408 nm), $^4I_{15/2}$–$^4F_{3/2}$ (446 nm),$^4I_{15/2}$–$^4F_{5/2}$ (453 nm),$^4I_{15/2}$–$^4F_{7/2}$ (491 nm), and $^4I_{15/2}$–$^2H_{11/2}$ (522 nm).The emission spectrum includes narrow bands, which are assigned to the following transitions: $^2H_{9/2}$–$^4I_{15/2}$ (409 nm), $^2H_{11/2}$–$^4I_{15/2}$ (524 nm), $^4S_{3/2}$–$^4I_{15/2}$ (549 nm), and $^4F_{9/2}$–$^4I_{15/2}$ (670 nm).The spectral line positions of $La_2O_2S:Er^{3+}$ spectra are the same. Change of host leads to the intensity redistribution, which is most pronounced for $^4I_{15/2}$–$^4G_{11/2}$ transition in the excitation spectrum.

Fig. 6 displays excitation and emission spectra of RE_2O_2S (RE = Y, La) activated by 1% of Eu^{3+} ions. The excitation spectrum of $Y_2O_2S:Eu^{3+}$ monitored at 545 nm ($^5D_1$–$^7F_1$) consists of following transitions:$^7F_0$–$^5D_4$ (353 nm),$^7F_0$–$^5L_7$(378 nm) and $^7F_2$–$^5D_2$(488 nm). The emission spectrum shows narrow bands originating from $^5D_1$and $^5D_0$ excited levels. Surprisingly, that the emission spectrum of $Y_2O_2S:Eu^{3+}$ is dominated by transition$^5D_1$–$^7F_1$(545 nm), whereas the most prominent luminescence bands are usually attributed to the $^5D_0$–^7F_J transition [46-48]. Such behavior was previously reported for $La_2O_2S:Eu^{3+}$ bulk phosphors [49, 50].

Dominance of $^5D_1$ emission can be explained by small phonon energy in regarded host, because significant amount of ionsrelax to $^5D_1$ level after the UV excitation , and they radiatively decay to the ground state before nonradiative decay to $^5D_0$ metastable level. We also observed $^5D_2$–$^7F_2$(490 nm), $^5D_1$–$^7F_3$ (587 nm), $^5D_0$–$^7F_1$ (592 nm), $^5D_0$–$^7F_2$ (621 nm) and $^5D_0$–$^7F_3$ (670 nm) transitions.

The excitation and emission spectra of $La_2O_2S:Eu^{3+}$ display situation, which is more usual for Eu^{3+}-doped compounds. The excitation spectrum of $Y_2O_2S:Eu^{3+}$ monitored at 623 nm ($^5D_0$–$^7F_2$) consists of broad intense band corresponding to charge transfer S^{2-}–Eu^{3+} (338 nm) and low-intense line assigned to the typical intra-configurational transitionsof the Eu^{3+} ion: $^7F_0$–$^5L_6$ (394 nm), $^7F_0$–$^5D_2$ (466 nm), $^7F_0$–$^5D_1$ (536 nm) and $^7F_1$–$^5D_0$ (593 nm). The emission spectrum is dominated by the forced electric dipole transition $^5D_0$–$^7F_2$ with maximum at 623 nm. Other observed lines are attributed to the $^5D_1$–$^7F_1$ (538 nm), $^5D_1$–$^7F_2$ (555 nm), $^5D_1$–$^7F_3$ (586 nm), $^5D_0$–$^7F_1$ (594 nm), $^5D_0$–$^7F_3$ (670 nm) and $^5D_0$–$^7F_4$ (704 nm).

Due to the unique luminescence properties of Eu^{3+} ions, it is quite easy to analyze the luminescent center local surrounding and its symmetry using only emission spectrum. The asymmetry ratio (R_{Eu}) gives information about local changes around the Eu^{3+} ions. It is defined as intensity ratio of forced electric dipole $^5D_0$–$^7F_2$ and magnetic dipole $^5D_0$–$^7F_1$ transitions. The higher the asymmetry parameter R_{Eu} is, the more apart from a centrosymmetric geometry luminescent center is located. The calculated R_{Eu} values of $Y_2O_2S:Eu^{3+}$ and $La_2O_2S:Eu^{3+}$ samples are 0.58 and 3.16, respectively. It is worth noting that the calculated R_{Eu} values of $Y_2O_2S:Eu^{3+}$ and $La_2O_2S:Eu^{3+}$ samples significantly differ, which indicate big difference in local surrounding of Eu^{3+} ions in these hosts.

The steady state luminescence spectra of RE_2O_2S (RE = Y, La) powders doped with 1% of Ce^{3+} ions are shown in Fig. 7. The excitation spectrum of $Y_2O_2S:Ce^{3+}$ sample displays two broad bands centered at 265 and 407 nm (λ_{em} = 545 nm). These bands correspond to direct excitation of the Ce^{3+} ions via transitions to the components of Ce^{3+} 5d configuration. The emission spectrum also consists of two lines attributed to allowed 5d–4f transition of Ce^{3+} ion. Generally, emission lines attributed to allowed 5d–4f transition in Ce^{3+}-doped materials are quite broad [51, 52]. Sometimes they are split into two components separated by approximately 2000 cm^{-1} due to the spin-orbit splitting of the $4f^1$ ground state into two components $^2F_{5/2}$ and $^2F_{7/2}$. The bands observed in $Y_2O_2S:Ce^{3+}$ and $La_2O_2S:Ce^{3+}$ exhibit splitting by 5000 cm^{-1} and cannot be due to splitting of the ground stated mentioned above. Therefore, two bands ib Ce^{3+} luminescence must be ascribed to the electron transitions from the lowest and second 5d levels to the ground state of Ce^{3+}[53].

Change of host to La_2O_2S does not affect spectroscopic properties of Ce^{3+}-doped material. The line positions are almost the same for both excitation and emission spectra.

4. Conclusions

In summary, a method for the production of luminescent materials on the basis of rare-earth oxysulfides was developed. The advantage of the method consists in the precipitation of a sulfur-containing precursor from a homogeneous nitrate solution and subsequent transformation in a reducing and sulfidating atmosphere. The use of chemical homogenization made it possible to achieve an excellent uniform distribution of cations in the structure. The use of sulfates as precursors, in view of the presence of sulfur in the structure, greatly simplifies the process of obtaining solid solutions of oxysulfides. All synthesized samples have single phase without any impurities. The excitation and emission spectra of RE_2O_2S:Ln (RE = Y, La; Ln = Dy, Er, Eu) consist of characteristic bands corresponding to the 4f-4f intra configurational transitions. The study of Dy^{3+} and Eu^{3+}-doped powders revealed that Y_2O_2S host possesses higher local symmetry than La_2O_2S one. The excitation and emission spectra of $RE_2O_2S:Ce^{3+}$ (RE = Y, La) phosphor displayed allowed 5d–4f transition.

Acknowledgement

The authors would like to thank the staff of the Engineering Center of the Tyumen State University for their help in carrying out physical and chemical tests. We also thank Andrey Bobylev for conducting electron microscopy.

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Captions

Fig.1. Experimental, calculated, and difference Rietveld plot of: a,b) $(RE_{0.99}Ln_{0.01})_2(SO_4)_3$; c,d) $(RE_{0.99}Ln_{0.01})_2(SO_4)_3$-$(RE_{0.99}Ln_{0.01})_2O_2SO_4$-$(RE_{0.99}Ln_{0.01})_2O_2S$-$(RE_{0.99}Ln_{0.01})_2O_3$; e,f) $(RE_{0.99}Ln_{0.01})_2O_2S$ - $(RE_{0.99}Ln_{0.01})_2O_3$; g,h) $(RE_{0.99}Ln_{0.01})_2O_2S$.

Fig. 2. SEM image of a) $(La_{0.99}Dy_{0.01})_2(SO_4)_3$; b) $(La_{0.99}Dy_{0.01})_2O_2S$

Fig. 3. Coordination polyhedron structure La_2O_2S

Fig. 4. Excitation and emission spectra of $Y_2O_2S:Dy^{3+}$and $La_2O_2S:Dy^{3+}$phosphors

Fig. 5. Excitation and emission spectra of $Y_2O_2S:Er^{3+}$ and $La_2O_2S:Er^{3+}$phosphors

Fig. 6. Excitation and emission spectra of $Y_2O_2S:Eu^{3+}$ and $La_2O_2S:Eu^{3+}$phosphors

Fig. 7. Excitation and emission spectra of $Y_2O_2S:Ce^{3+}$ and $La_2O_2S:Ce^{3+}$phosphors

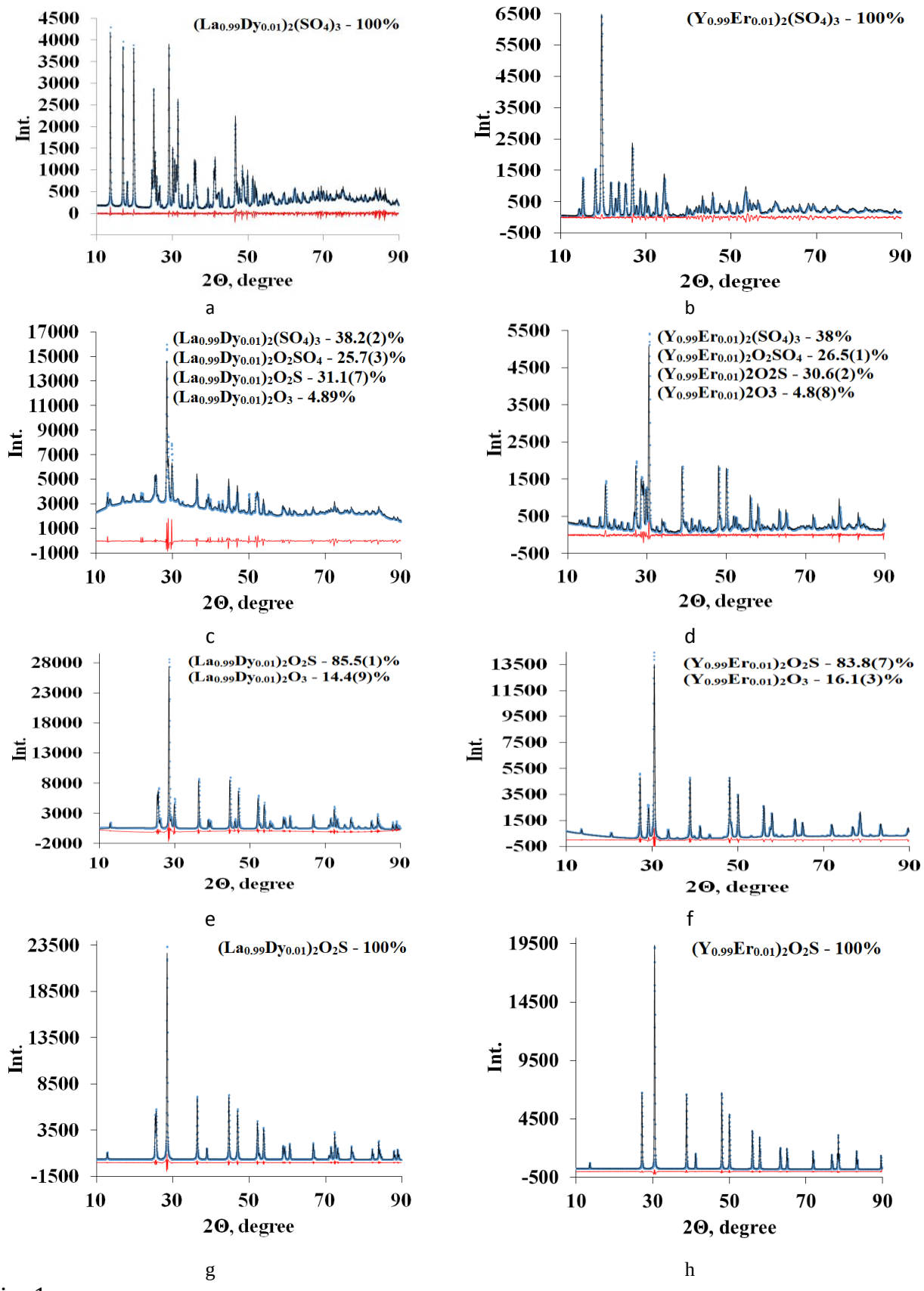
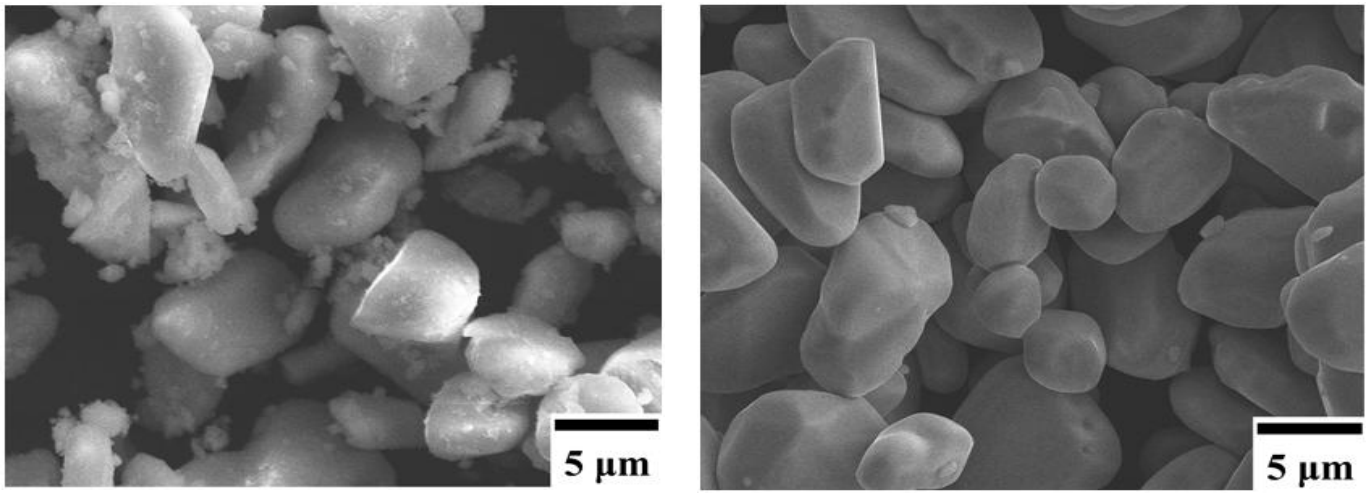

(La0.99Dy0.01)2(SO4)3 - 100%
Int.
2Θ, degree
a
(Y0.99Er0.01)2(SO4)3 - 100%
Int.
2Θ, degree
b
(La0.99Dy0.01)2(SO4)3 - 38.2(2)%
(La0.99Dy0.01)2O2SO4 - 25.7(3)%
(La0.99Dy0.01)2O2S - 31.1(7)%
(La0.99Dy0.01)2O3 - 4.89%
Int.
2Θ, degree
c
(Y0.99Er0.01)2(SO4)3 - 38%
(Y0.99Er0.01)2O2SO4 - 26.5(1)%
(Y0.99Er0.01)2O2S - 30.6(2)%
(Y0.99Er0.01)2O3 - 4.8(8)%
Int.
2Θ, degree
d
(La0.99Dy0.01)2O2S - 85.5(1)%
(La0.99Dy0.01)2O3 - 14.4(9)%
Int.
2Θ, degree
e
(Y0.99Er0.01)2O2S - 83.8(7)%
(Y0.99Er0.01)2O3 - 16.1(3)%
Int.
2Θ, degree
f
(La0.99Dy0.01)2O2S - 100%
Int.
2Θ, degree
g
(Y0.99Er0.01)2O2S - 100%
Int.
2Θ, degree
h

Fig. 1



a) b)

Fig. 2

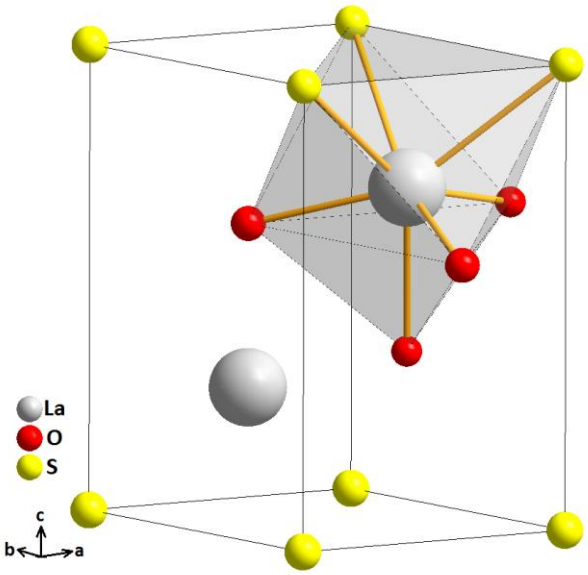


Fig. 3.

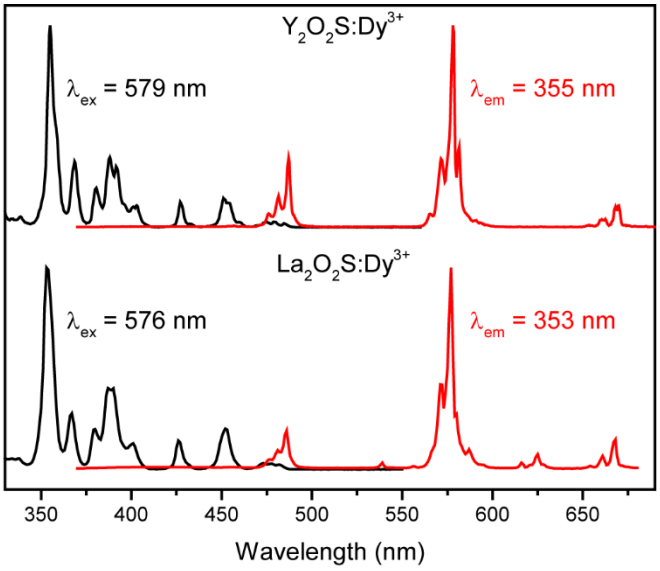


Fig. 4

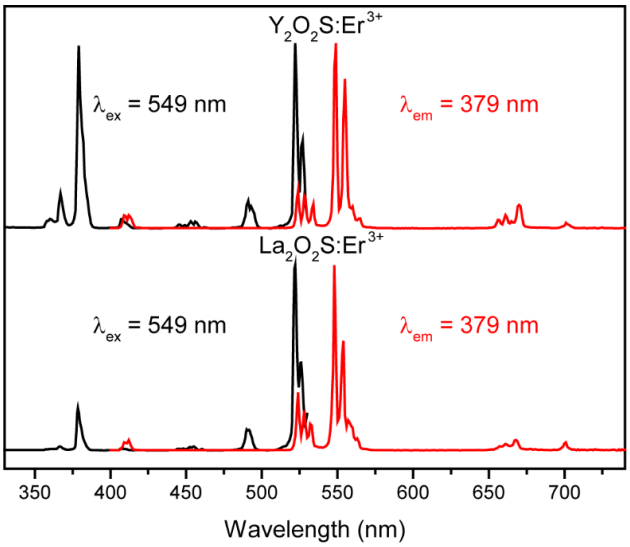


Fig. 5

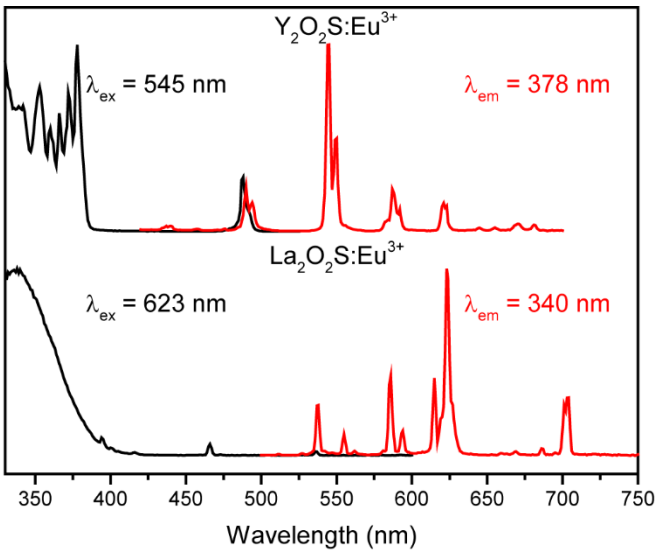


Fig. 6

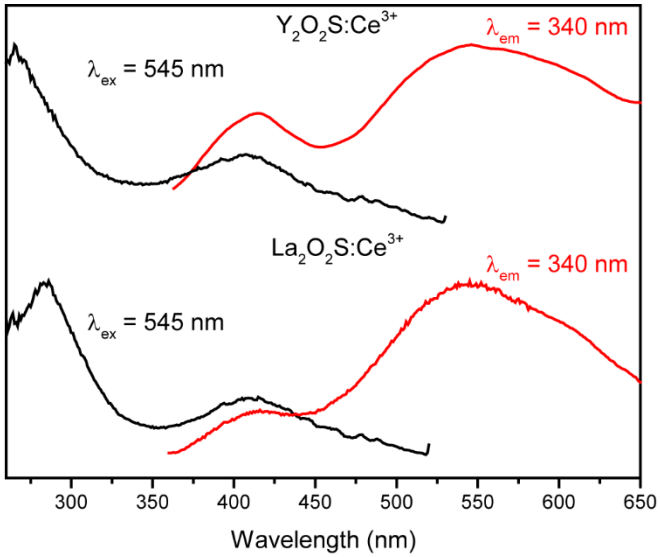

$Y_2O_2S:Ce^{3+}$
λ_{em} = 340 nm
λ_{ex} = 545 nm
$La_2O_2S:Ce^{3+}$
λ_{em} = 340 nm
λ_{ex} = 545 nm
300
350
400
450
500
550
600
650
Wavelength (nm)

Fig. 7

Table(s)

Table 1. Main parameters of processing and refinement of the samples La-O-S

Compound	Phase	Weight (%)	Space group	Cell parametrs (°, Å), Cell volume (Å^3)	R_{wp}, R_p (%)	X^2
$(La_{0.99}Dy_{0.01})_2(SO_4)_3$	$(La_{0.99}Dy_{0.01})_2(SO_4)_3$	100	*C2/c*	a=22.1136 b=6.9036 c=7.1198 V=1035.15(3)	6.34 4.77	1.51
$(La_{0.99}Dy_{0.01})_2(SO_4)_3$ - $(La_{0.99}Dy_{0.01})_2O_2SO_4$ - $(La_{0.99}Dy_{0.01})_2O_2S$ - $(La_{0.99}Dy_{0.01})_2O_3$	$(La_{0.99}Dy_{0.01})_2(SO_4)_3$	38.2(2)	*C2/c*	a=22.1138 b=6.9026 c=7.1204 V=1035.18(1)	5.31 4.35	7.72
	$(La_{0.99}Dy_{0.01})_2O_2SO_4$	25.7(3)	*C2/c*	a=14.3493 b=4.2858 c=8.3882 V=493.43(1)		
	$(La_{0.99}Dy_{0.01})_2O_2S$	31.1(7)	*P-3m 1*	a=4.0520 c=6.9420 V=98.70(8)		
	$(La_{0.99}Dy_{0.01})_2O_3$	4.7(9)	*P-3m 1*	a=4.0520 c=6.9420 V=82.67(0)		
$(La_{0.99}Dy_{0.01})_2O_2S$ - $(La_{0.99}Dy_{0.01})_2O_3$	$(La_{0.99}Dy_{0.01})_2O_2S$	85.5(1)	*P-3m 1*	a=4.0521 c=6.9428 V=98.72(5)	12.97 9.91	9.12
	$(La_{0.99}Dy_{0.01})_2O_3$	14.4(9)	*P-3m 1*	a=3.9402 c=6.1493 V=82.67(9)		
$(La_{0.99}Dy_{0.01})_2O_2S$	$(La_{0.99}Dy_{0.01})_2O_2S$	100	*P-3m 1*	a=4.0520 c=6.9425 V=98.71(6)	4.71 3.51	1.46

Table 2. Main parameters of processing and refinement of the samples Y-O-S

Compound	Phase	Weight (%)	Space group	Cell parametrs (°, Å), Cell volume (Å^3)	R_{wp}, R_p (%)	X^2
$(Y_{0.99}Er_{0.01})_2(SO_4)_3$	$Y_{0.99}Er_{0.01})_2(SO_4)_3$	100	*Pbcn*	a=12.4770 b=9.0354 c=9.8348 V=1108.72(3)	10.68 7.67	3.22
$(Y_{0.99}Er_{0.01})_2(SO_4)_3$ - $(Y_{0.99}Er_{0.01})_2O_2SO_4$ - $(Y_{0.99}Er_{0.01})_2O_2S$ - $(Y_{0.99}Er_{0.01})_2O_3$	$(Y_{0.99}Er_{0.01})_2(SO_4)_3$	38.0(0)	*Pbcn*	a=12.4804 b=9.0406 c=9.8392 V=1110.16(0)	6.54 5.11	1.22
	$(Y_{0.99}Er_{0.01})_2O_2SO_4$	26.5(1)	*C2/c*	a=14.3439 b=4.2850 c=8.3858 V=493.09(1)		
	$(Y_{0.99}Er_{0.01})_2O_2S$	30.6(2)	*P-3m 1*	a=3.7800 c=6.5631 V=81.21(3)		
	$(Y_{0.99}Er_{0.01})_2O_3$	4.8(8)		a=10.6051 V=1192.73(6)		
$(Y_{0.99}Er_{0.01})_2O_2S$ - $(Y_{0.99}Er_{0.01})_2O_3$	$(Y_{0.99}Er_{0.01})_2O_2S$	83.8(7)	*P-3m 1*	a=3.7801 c=6.5638 V=81.22(6)	7.98 5.59	2.81
	$(Y_{0.99}Er_{0.01})_2O_3$	16.1(3)	*Ia-3*	a=10.6055 V=1192.87(1)		
$(Y_{0.99}Er_{0.01})_2O_2S$	$(Y_{0.99}Er_{0.01})_2O_2S$	100	*P-3m 1*	a=3.7800 c=6.5632 V=81.21(4)	5.01 3.31	1.06

Synthesis and Optical Properties RE_2O_2S:Ln (RE = La, Y; Ln = Ce, Eu, Dy, Er)

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Captions

Fig.1. Experimental, calculated, and difference Rietveld plot of: a,b) $(RE_{0.99}Ln_{0.01})_2(SO_4)_3$; c,d) $(RE_{0.99}Ln_{0.01})_2(SO_4)_3$-$(RE_{0.99}Ln_{0.01})_2O_2SO_4$-$(RE_{0.99}Ln_{0.01})_2O_2S$-$(RE_{0.99}Ln_{0.01})_2O_3$; e,f) $(RE_{0.99}Ln_{0.01})_2O_2S$ - $(RE_{0.99}Ln_{0.01})_2O_3$; g,h) $(RE_{0.99}Ln_{0.01})_2O_2S$.

Fig. 2. SEM image of a) $(La_{0.99}Dy_{0.01})_2(SO_4)_3$; b) $(La_{0.99}Dy_{0.01})_2O_2S$

Fig. 3. Coordination polyhedron structure La_2O_2S

Fig. 4. Excitation and emission spectra of $Y_2O_2S:Dy^{3+}$and $La_2O_2S:Dy^{3+}$phosphors

Fig. 5. Excitation and emission spectra of $Y_2O_2S:Er^{3+}$ and $La_2O_2S:Er^{3+}$phosphors

Fig. 6. Excitation and emission spectra of $Y_2O_2S:Eu^{3+}$ and $La_2O_2S:Eu^{3+}$phosphors

Fig. 7. Excitation and emission spectra of $Y_2O_2S:Ce^{3+}$ and $La_2O_2S:Ce^{3+}$phosphors

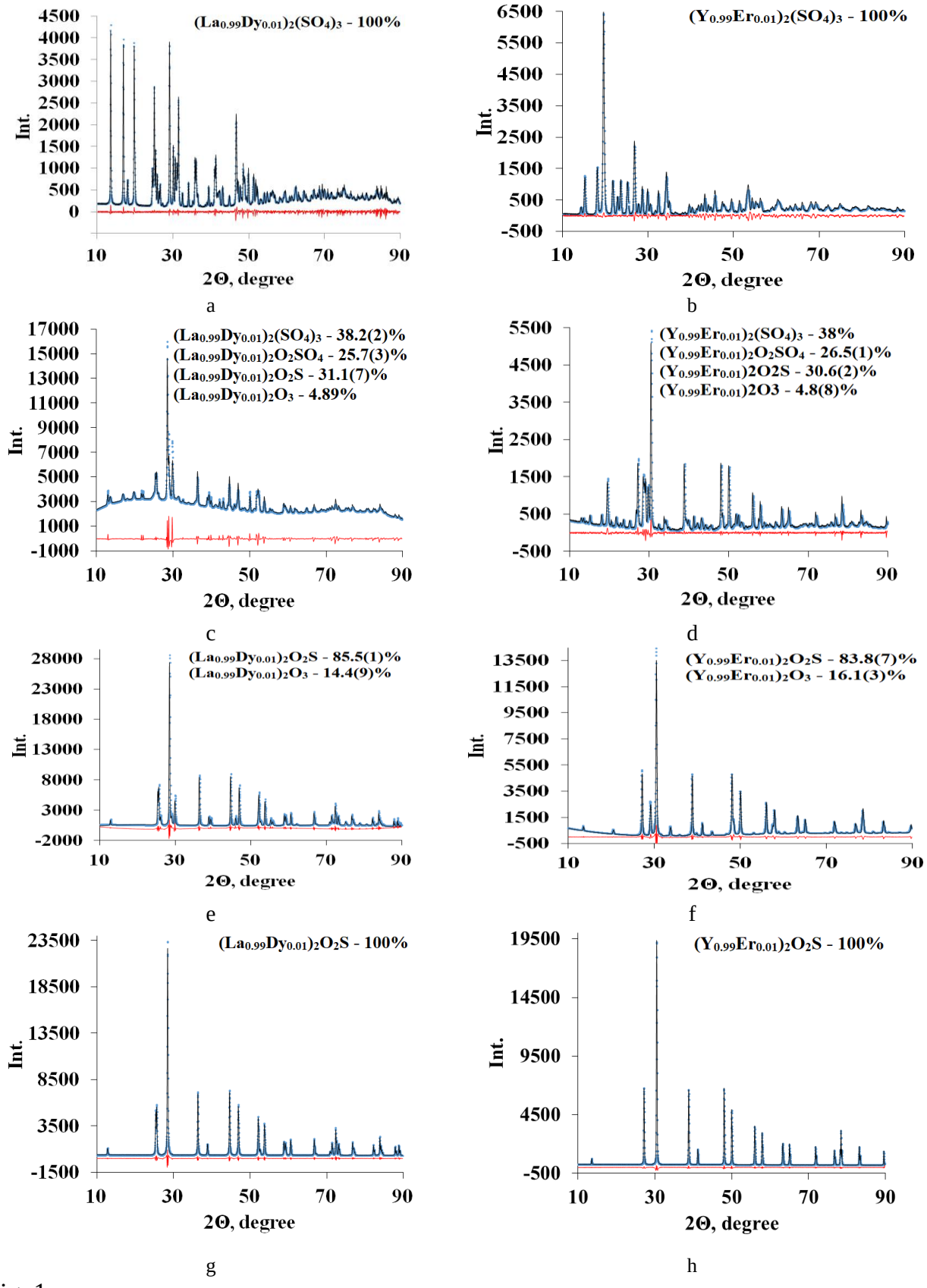
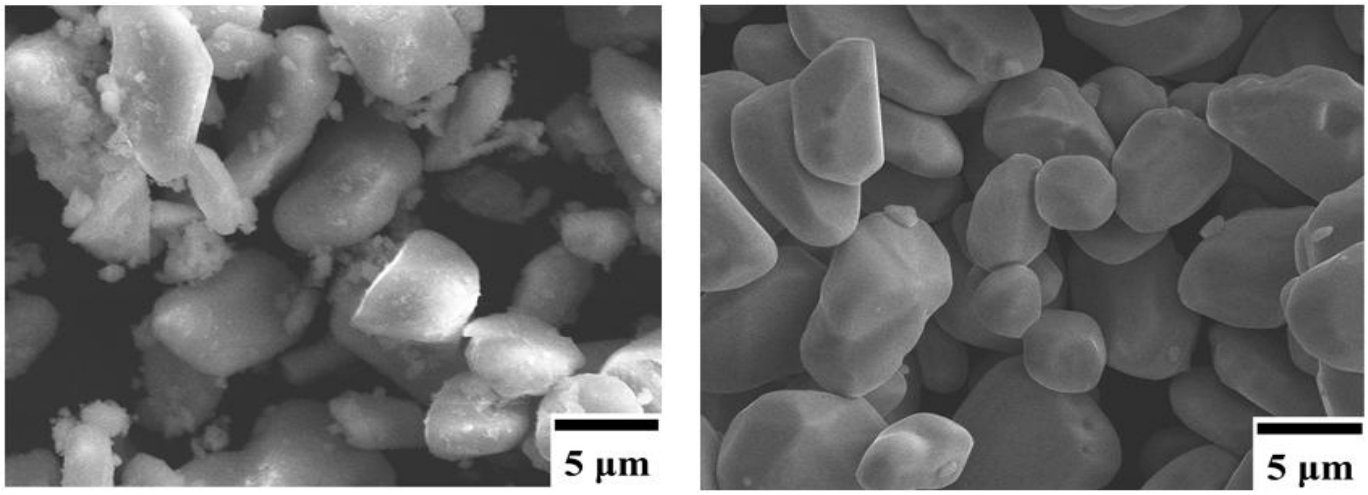

(La0.99Dy0.01)2(SO4)3 - 100%
Int.
2Θ, degree
a
(Y0.99Er0.01)2(SO4)3 - 100%
Int.
2Θ, degree
b
(La0.99Dy0.01)2(SO4)3 - 38.2(2)%
(La0.99Dy0.01)2O2SO4 - 25.7(3)%
(La0.99Dy0.01)2O2S - 31.1(7)%
(La0.99Dy0.01)2O3 - 4.89%
Int.
2Θ, degree
c
(Y0.99Er0.01)2(SO4)3 - 38%
(Y0.99Er0.01)2O2SO4 - 26.5(1)%
(Y0.99Er0.01)2O2S - 30.6(2)%
(Y0.99Er0.01)2O3 - 4.8(8)%
Int.
2Θ, degree
d
(La0.99Dy0.01)2O2S - 85.5(1)%
(La0.99Dy0.01)2O3 - 14.4(9)%
Int.
2Θ, degree
e
(Y0.99Er0.01)2O2S - 83.8(7)%
(Y0.99Er0.01)2O3 - 16.1(3)%
Int.
2Θ, degree
f
(La0.99Dy0.01)2O2S - 100%
Int.
2Θ, degree
g
(Y0.99Er0.01)2O2S - 100%
Int.
2Θ, degree
h

Fig. 1

1



a) b)

Fig. 2

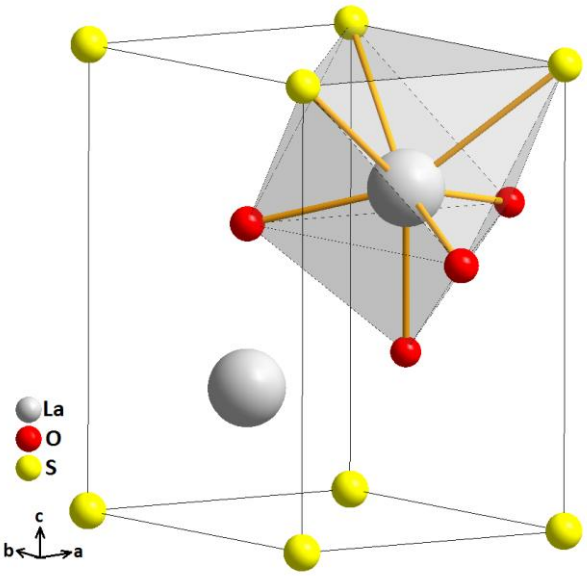


Fig. 3.

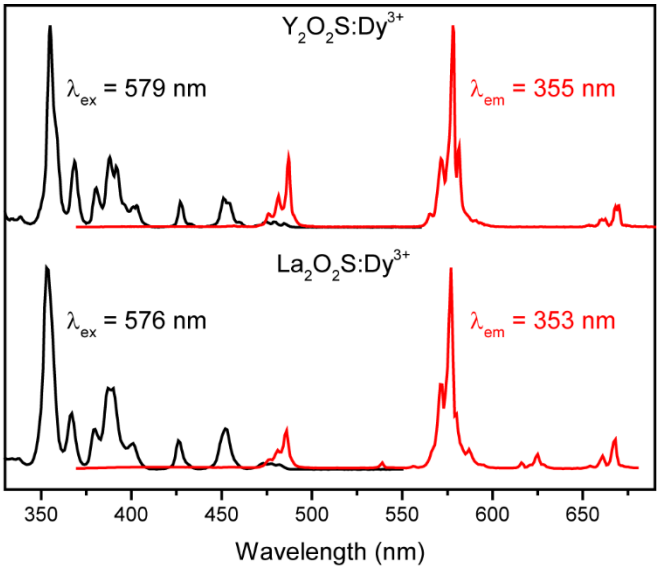
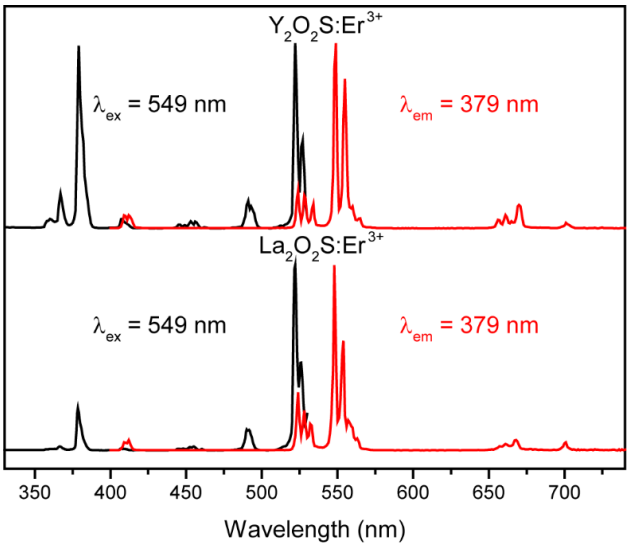


Fig. 4



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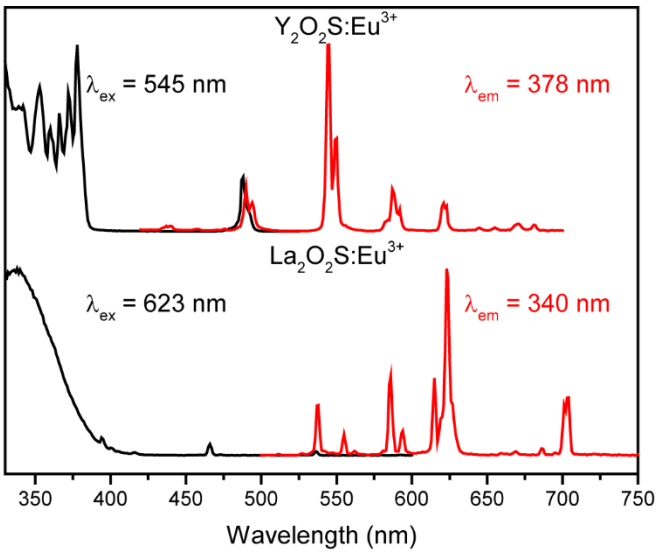

Fig. 6

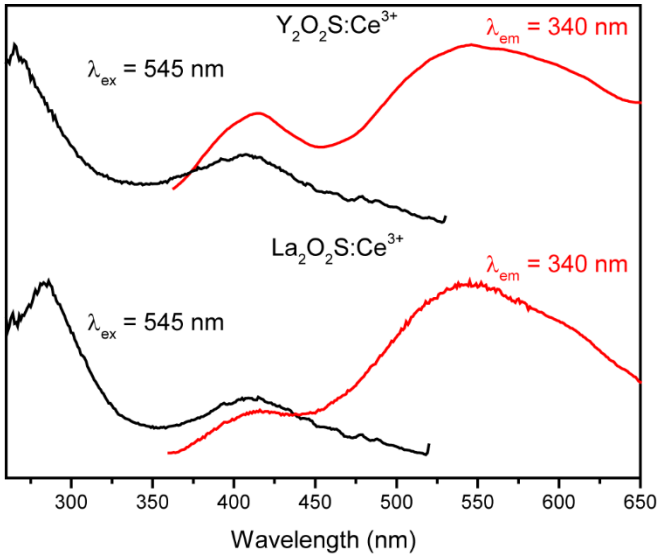
$Y_2O_2S:Ce^{3+}$
λ_{em} = 340 nm
λ_{ex} = 545 nm
$La_2O_2S:Ce^{3+}$
λ_{em} = 340 nm
λ_{ex} = 545 nm
300
350
400
450
500
550
600
650
Wavelength (nm)

Fig. 7

Figure(s)

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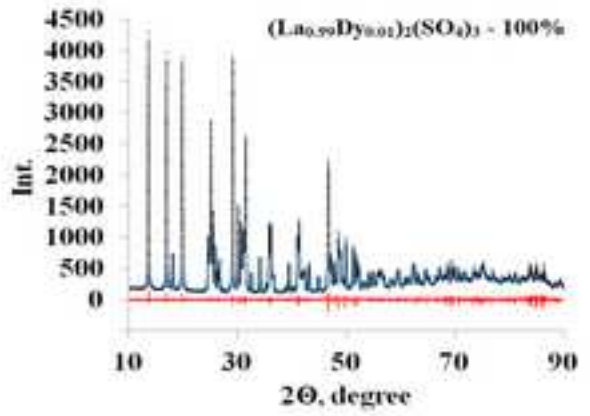

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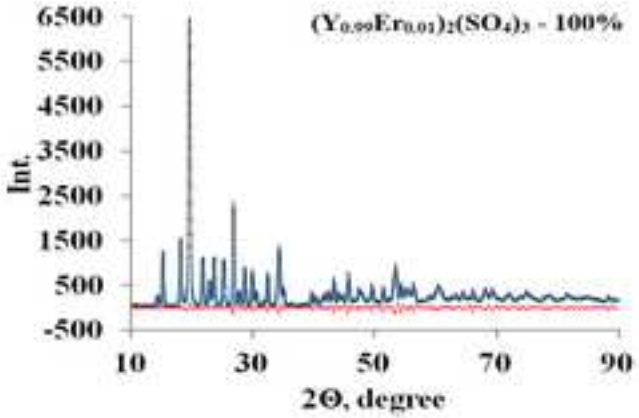

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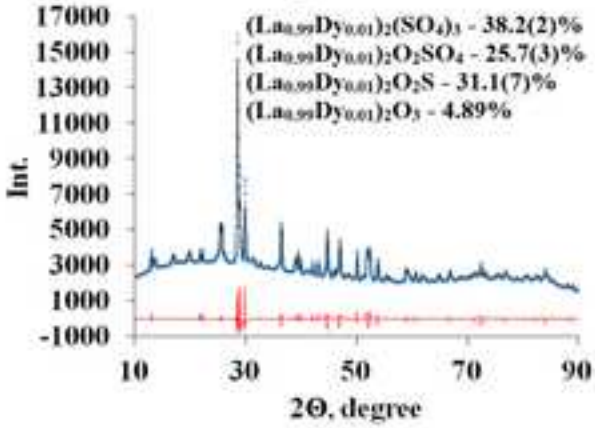

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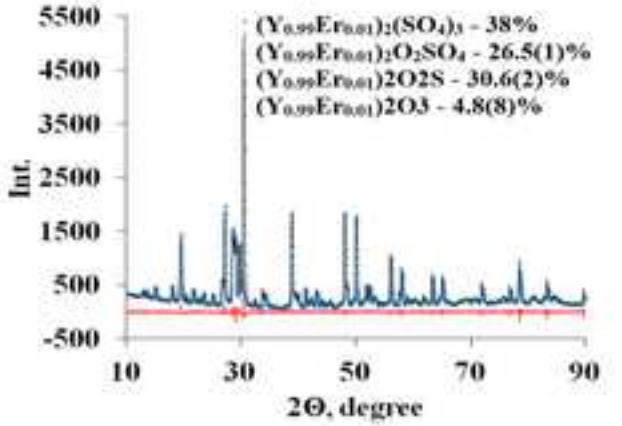

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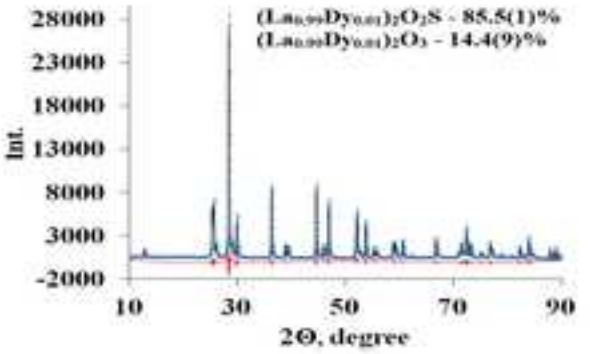

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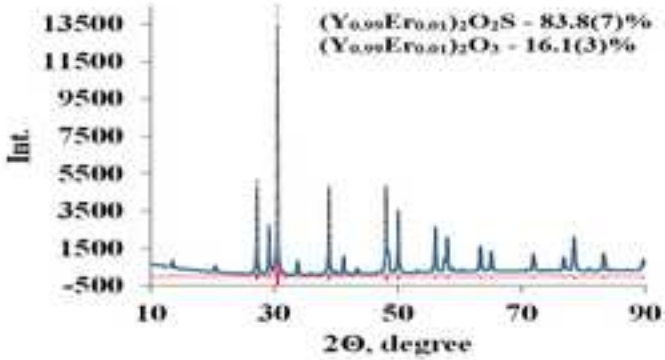

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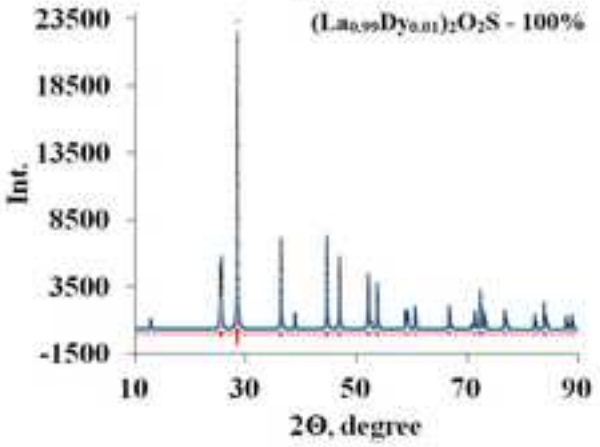

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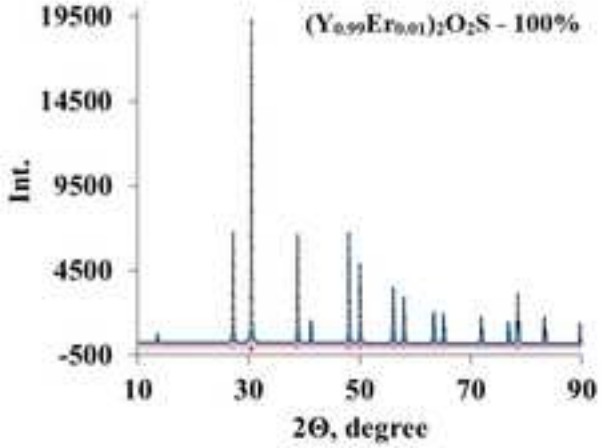

h

Figure(s)
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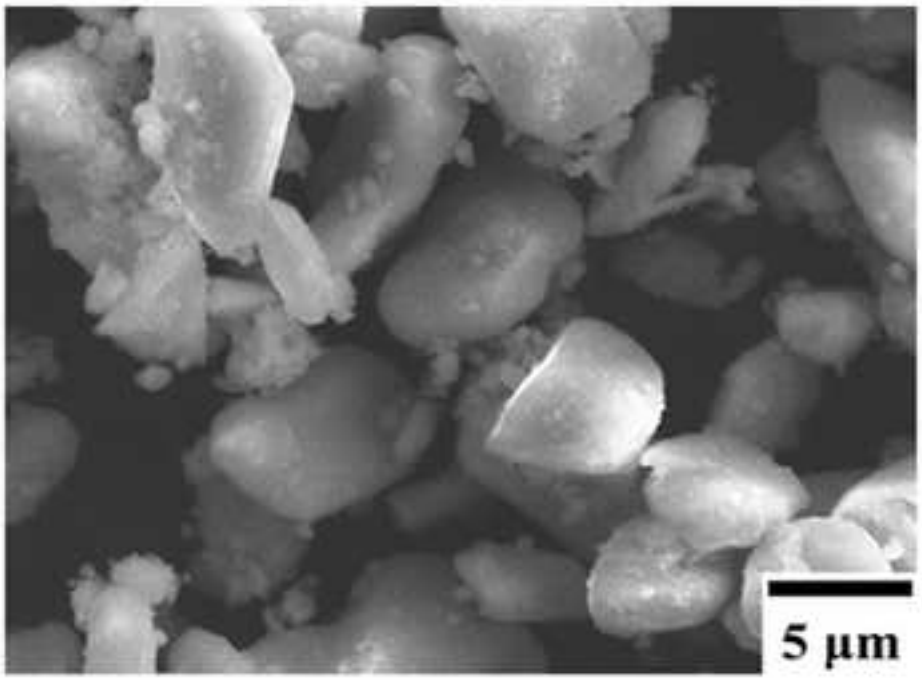

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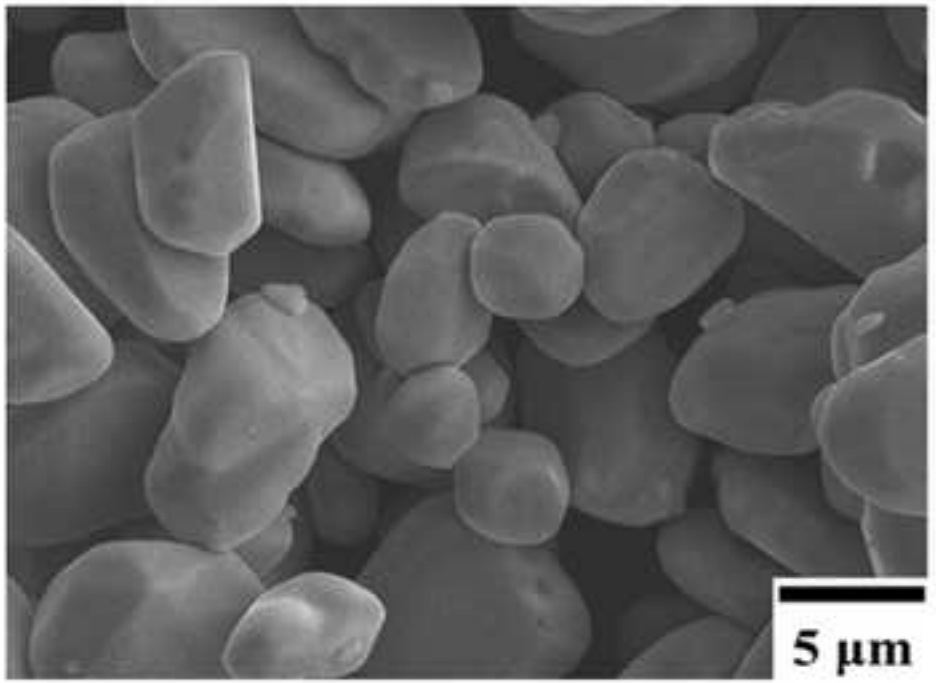

b)

Figure(s)

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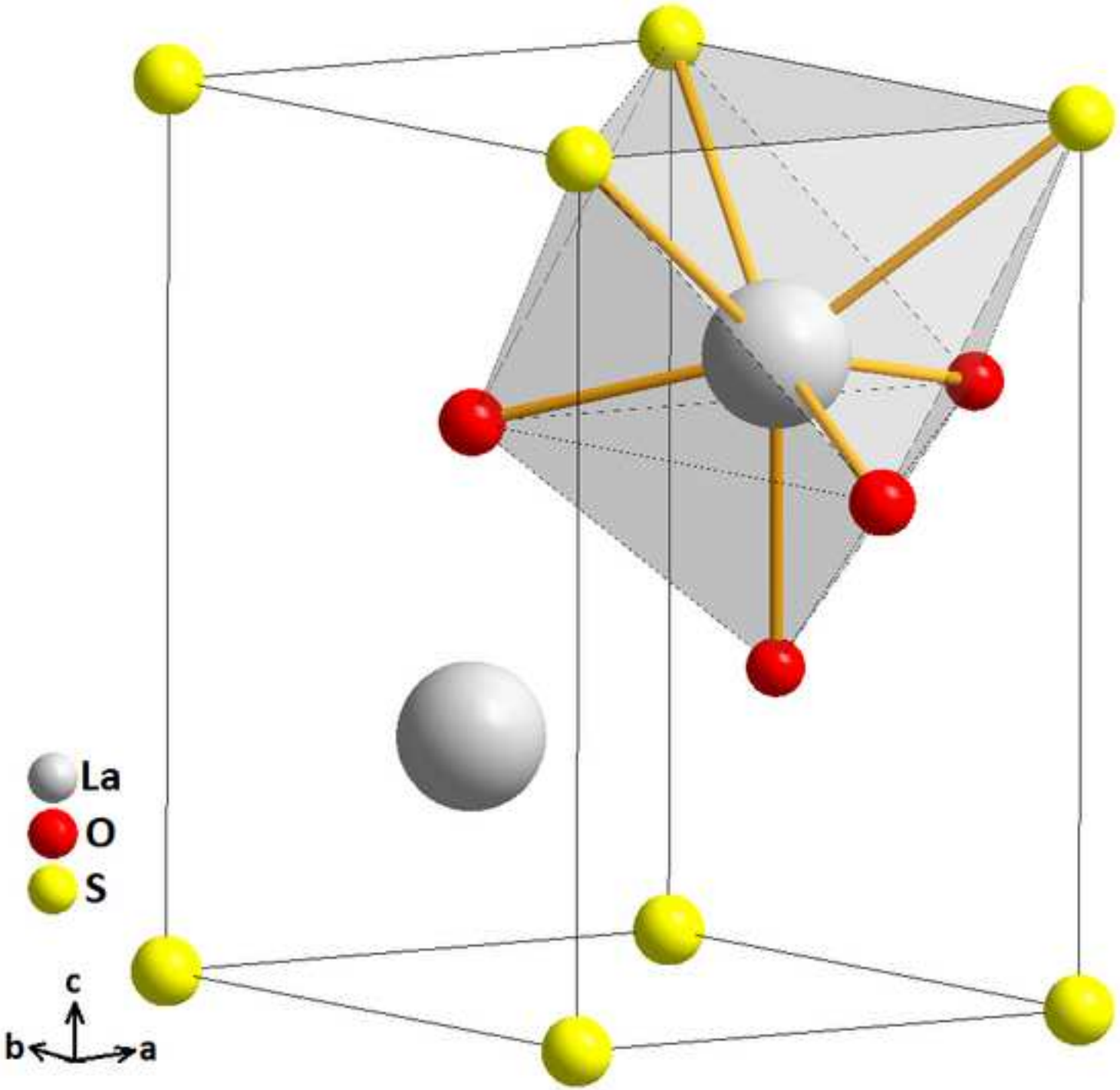

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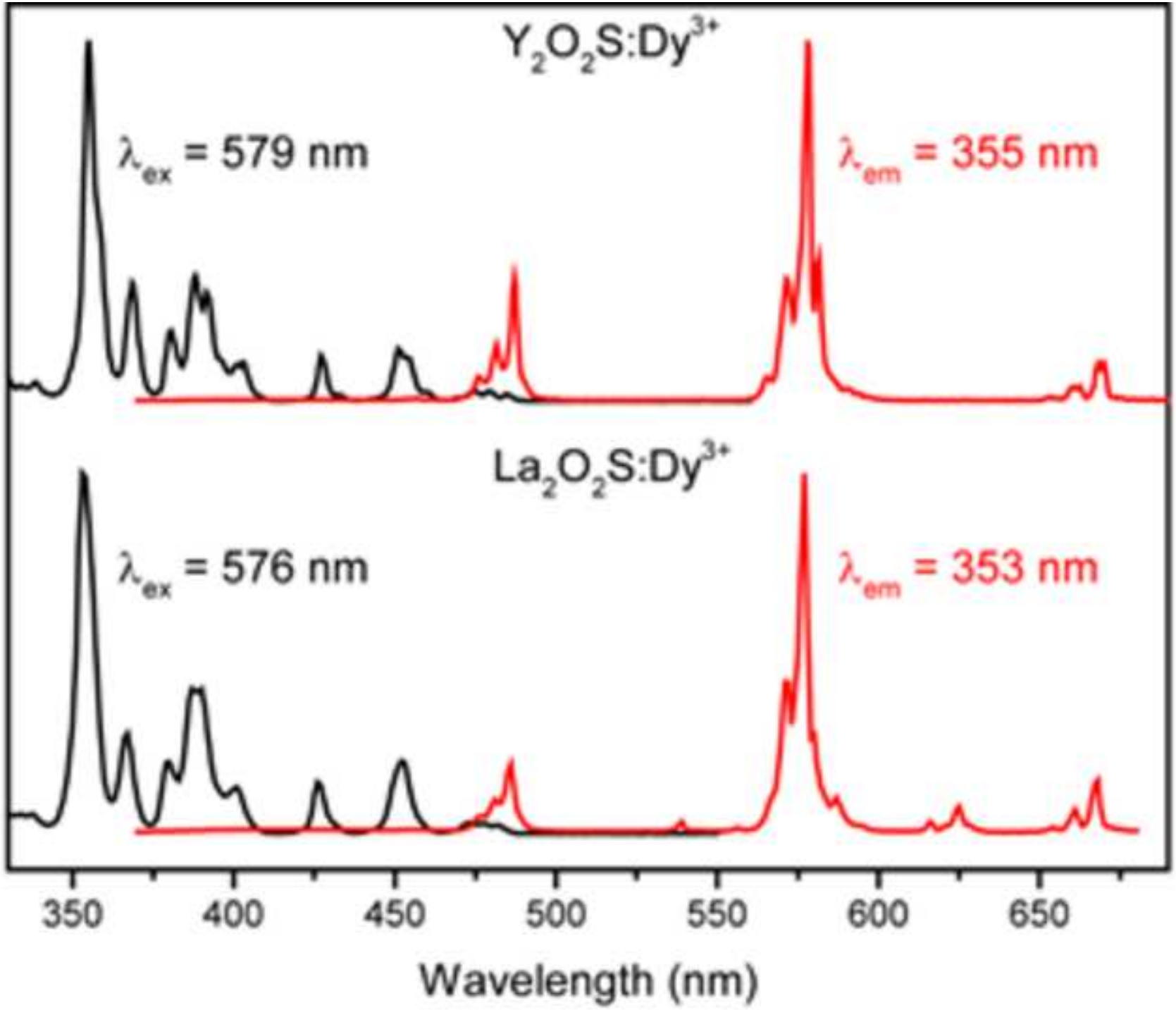

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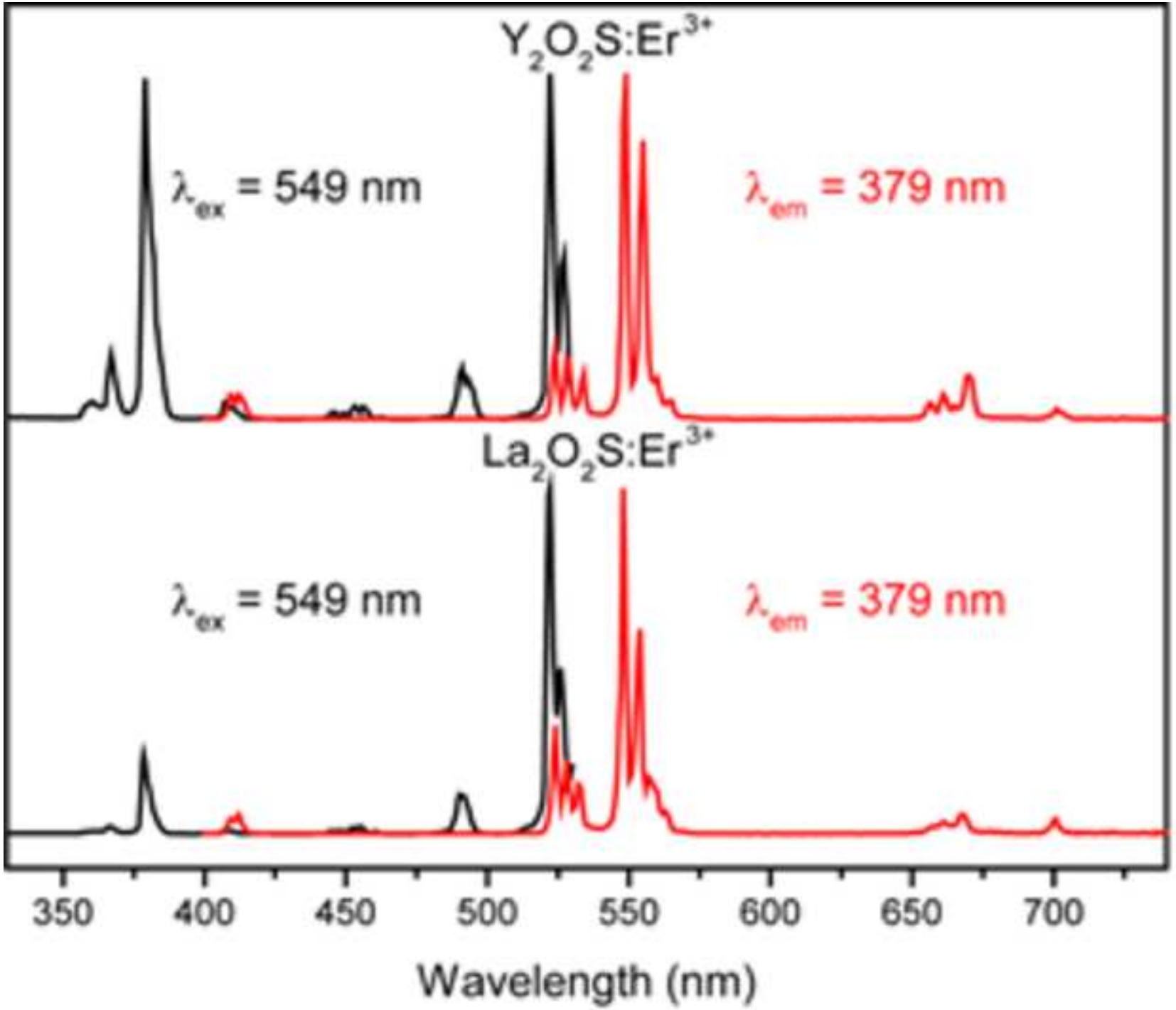

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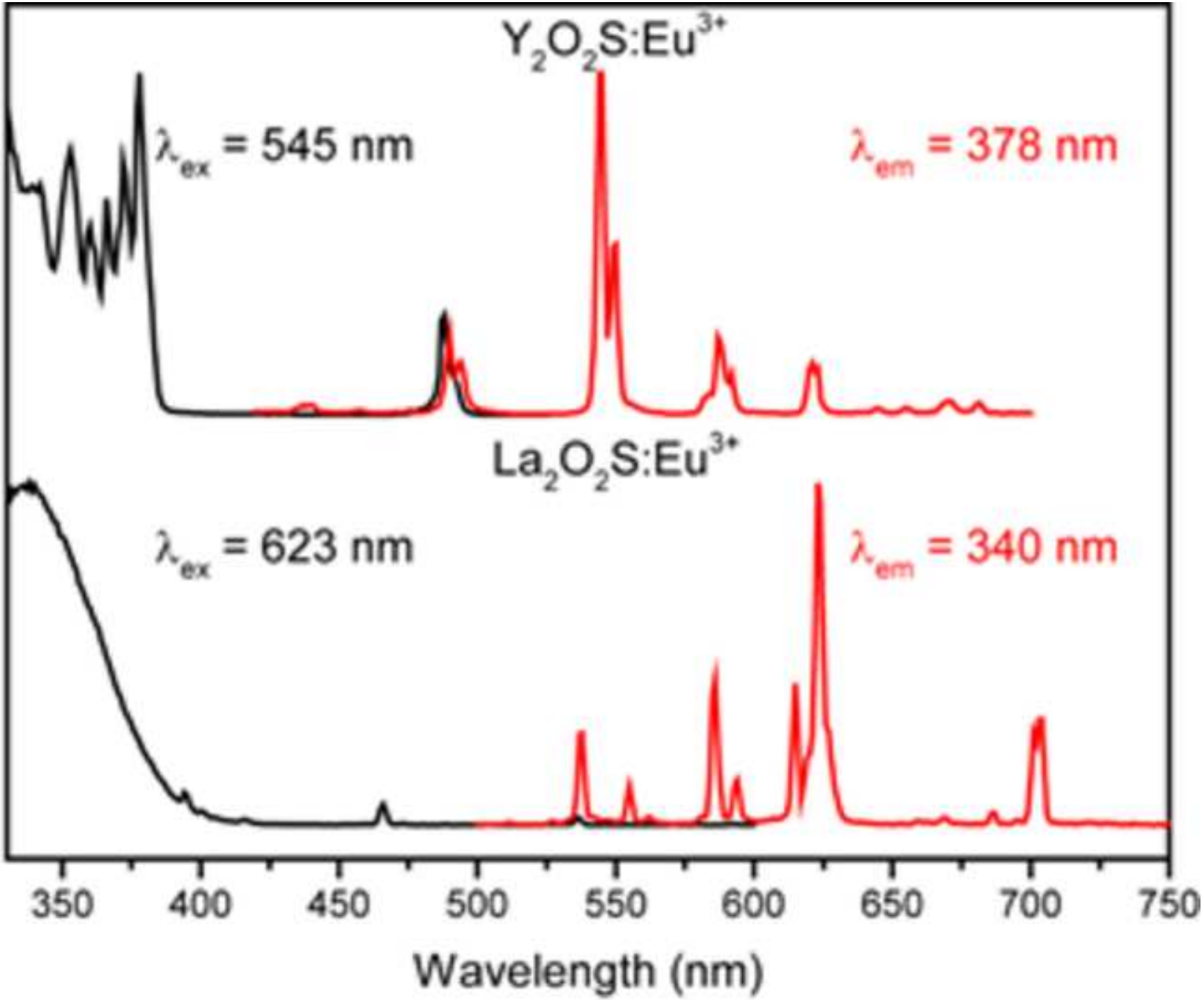

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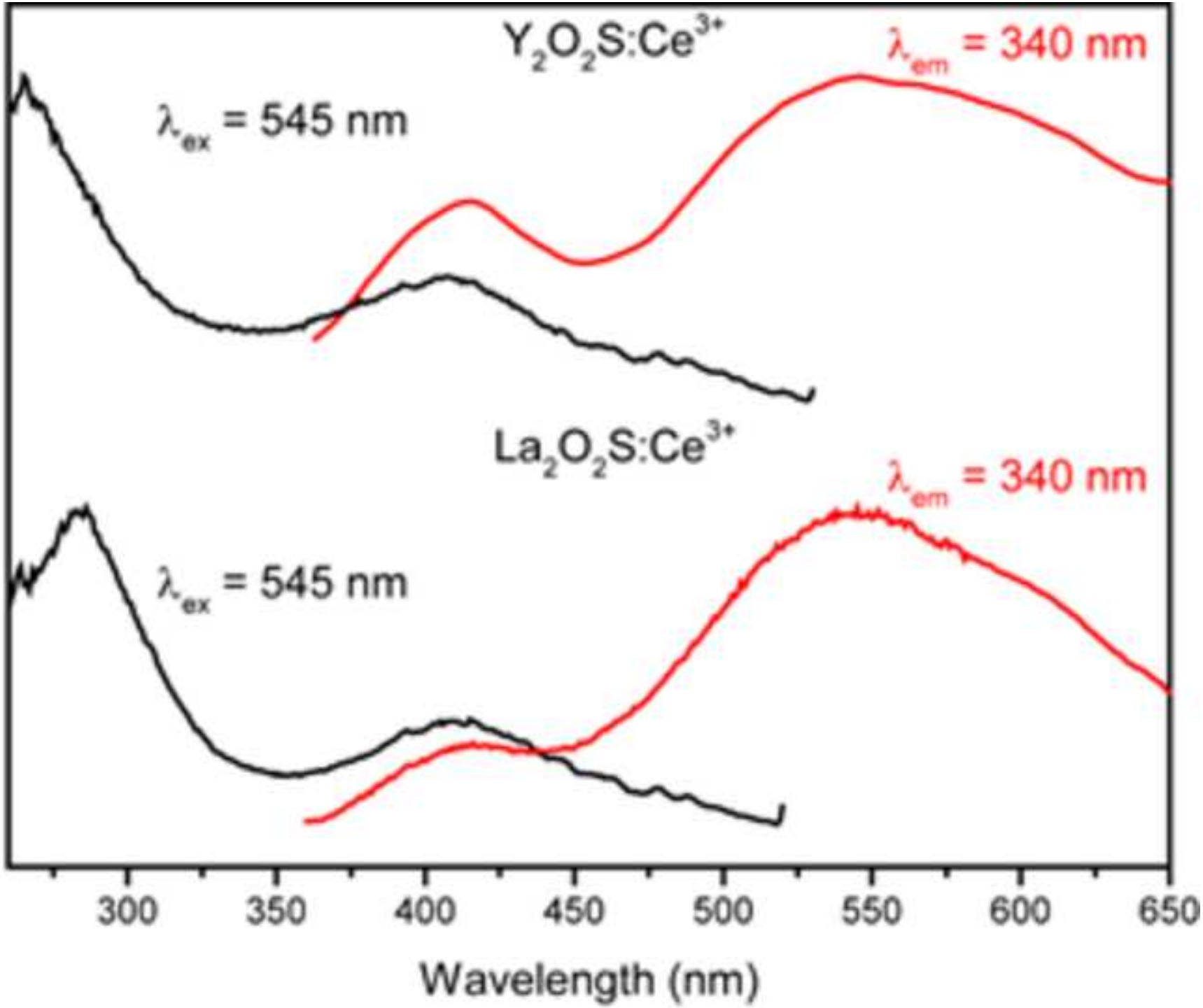

*Graphical Abstract

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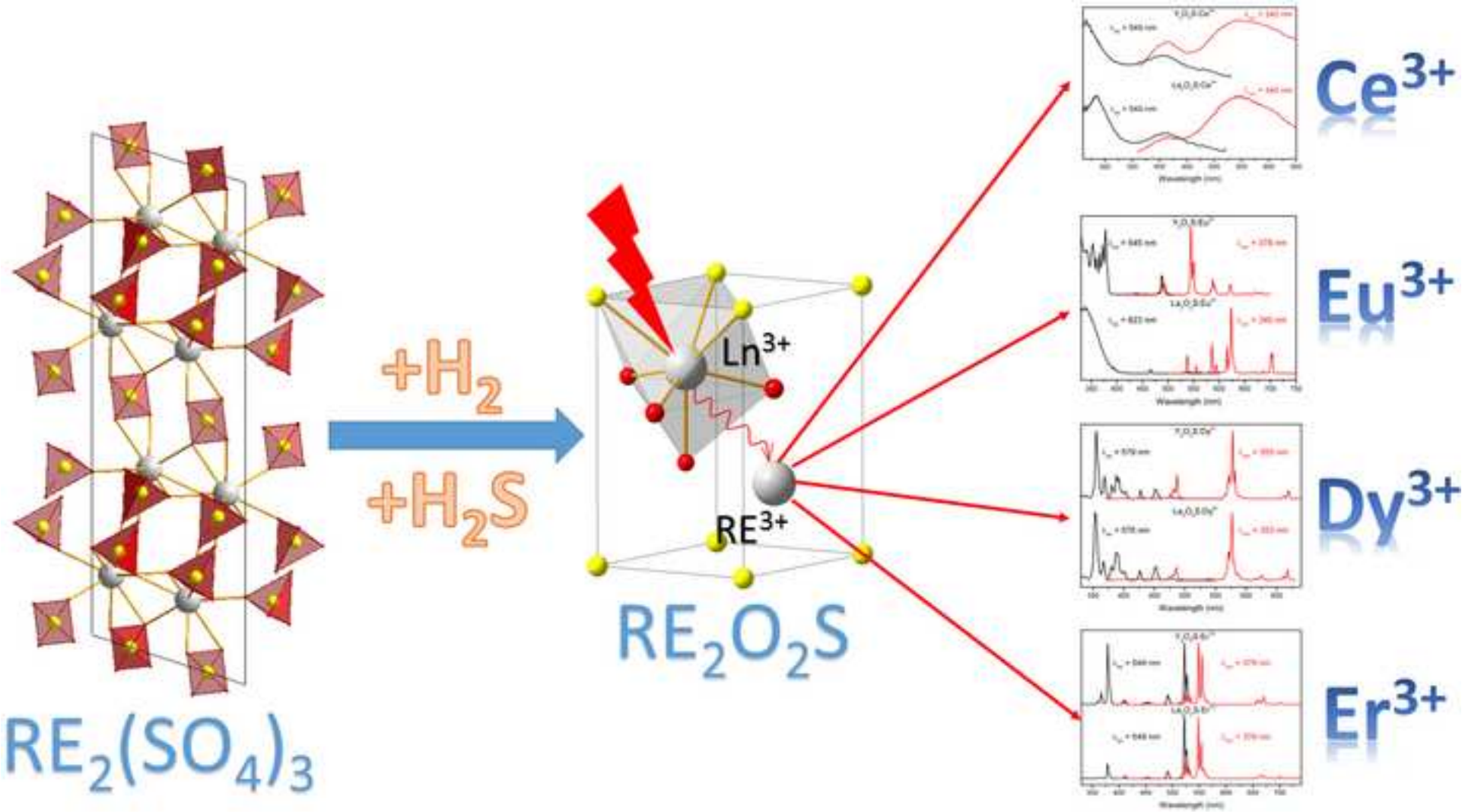

*Graphical Abstract Legend (TOC Figure)

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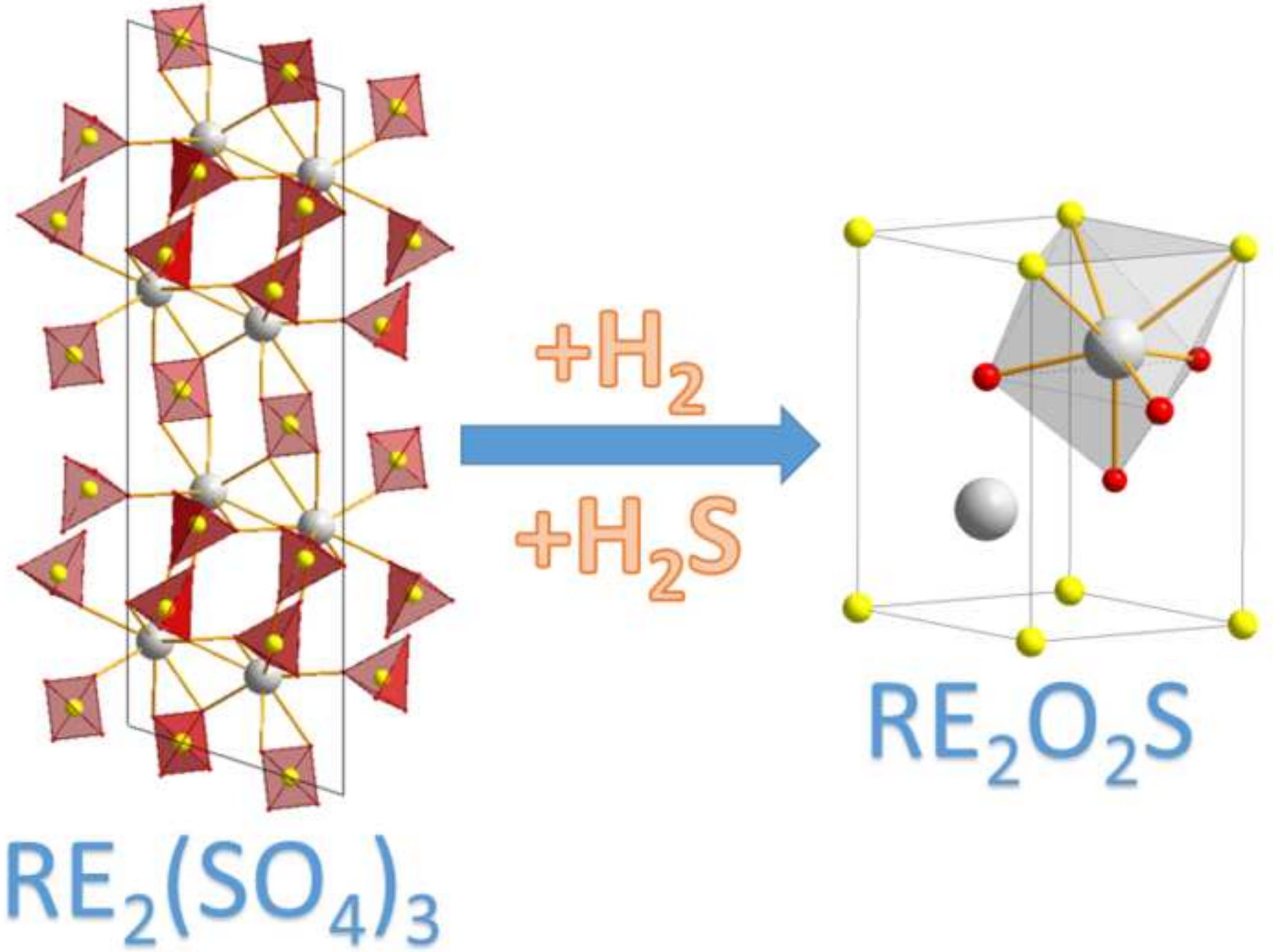

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