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Structural and spectroscopic properties of new noncentrosymmetric self-activated borate $Rb_3EuB_6O_{12}$ with B_5O_{10} units

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Abstract

New noncentrosymmetric double borate $Rb_3EuB_6O_{12}$ was synthesized by solid state reaction method, and their crystallographic parameters were obtained by Rietveld analysis. This borate crystallizes in the trigonal space group *R*32 with cell parameters of a = 13.4604(2) Å, c = 30.7981(5) Å, Z = 15. Their structure features a three-dimensional framework composed of $[B_5O_{10}]^{5-}$ groups that are bridged by Eu-O polyhedra. The existence of B_5O_{10} group in the structure was confirmed by vibrational spectroscopy. $Rb_3EuB_6O_{12}$ melts incongruently at 1101 K. The second harmonic generation effect of $Rb_3EuB_6O_{12}$ is 16 times higher than that of α-quartz standard. In the luminescence spectrum, the domination of a single prominent narrow line from hypersensitive $^5D_0$ - $^7F_2$ manifold of Eu^{3+} ions is observed, while $^5D_0$ - $^7F_1$ manifold and ultranarrow $^5D_0$ - $^7F_0$ line are of comparable peak intensity. These features are explained by specific local symmetry of Eu ion within the crystal structure of $Rb_3EuB_6O_{12}$..

Keywords: Rubidium rare earth borate; Solid state reaction; Rietveld refinement; Raman; IR, luminescence/

1. Introduction

During last decades, borate crystals are extensively studied mainly in connection with valuable properties of several nonlinear optical (NLO) borate crystals discovered in the past. Among the most attractive features of borate crystals such key characteristics as their transparency in deep UV range, high optical damage threshold combined with appropriate nonlinearity, and suitable chemical stability of optical surface should be pointed, and they define wide applications of borate materials in modern laser optics and nanophotonics [1-11]. Within borate matrices, boron ions can be either three- or four-fold coordinated, and the combination of BO_4 and BO_3 units produces large variety of crystal structures with differing properties [12-15]. In recent years, many complex borate crystals were discovered and evaluated in searching for new efficient polyfunctional materials for nonlinear optics and photonics [16-23].

One of extensively studying trends concerning borates' crystals is doping them with rare earth ions producing either phosphors or lasing crystals with the possibility of simultaneous frequency doubling of generated radiation within the same crystal. Such noncentrosymmetric (NCS) borates as $YAl_3(BO_3)_4$ (YAB) [24] $Ca_4RO(BO_3)_3$ (R = rare earth of Y) [25] and $Na_3La_9O_3(BO_3)_8$ [26] are known as self-frequency doubling (SFD) materials when doped with Nd. $YAl_3(BO_3)_3$ doped with Mn^{4+} or pair Yb^{3+}/Tm^{3+} is prospective as efficient phosphor material [27,28]. Femtosecond pulse generation was carried out in $Ca_3Gd_2(BO_3)_4:Nd^{3+}$ crystal [29]. However, within $M_2O–RE_2O_3–B_2O_3$ (M = alkali metal, RE = rare earth) systems, only several NCS crystalline media have been found up to date, namely $Na_3La_2(BO_3)_3$ [30], $Rb_3Y_2B_3O_9$ [31] $Na_3La_9O_3(BO_3)_8$ [32], $K_3REB_6O_{12}$ (RE = Y, Eu, Tb, Lu) [33-36] and $Rb_3NdB_6O_{12}$ [37]. It is worth noting that $K_3REB_6O_{12}$ (RE = Y, Eu, Tb, Lu) and $Rb_3NdB_6O_{12}$ are related to the same structure type, and they crystallize in the trigonal space group *R*32 and exhibit noticeable SHG signal. The UV cut-off edge of $K_3YB_6O_{12}$ and $Rb_3NdB_6O_{12}$ is reported to be at 195-197 nm [33,37]. These findings indicate that the $M_2O–RE_2O_3–B_2O_3$ ternary systems are complex enough, and other NCS borate crystals transparent in UV range can be formed. In the present

study, we employ larger alkali ions since they could have stronger polarizability and, hence, stronger overall NLO response can be expected. The structural and optical properties of $Rb_3Y_2B_3O_9$ are not investigated up to date. Recently, we reported the properties of new noncentrosymmetric borate crystal, $Rb_3NdB_6O_{12}$ [37], which is featured by unusual blue shift if luminescent lines with respect to traditional laser materials. This shift is larger than that in other borate family, namely, huntites, and it is close to that observed in fluorides. The observed peculiarity must be ascribed to specific influence of the current boron-based framework structure. Respectively, it is interesting to examine the properties of other rare earth ions in the same structure. In the present study, the synthesis of new crystal $Rb_3EuB_6O_{12}$ is performed, its structural and vibrational characteristics are investigated, and luminescent properties of Eu^{3+} ion in this new borate structure are studied.

2. Experimental

The samples were prepared by solid state reactions using high-purity starting reagents (Red Chemist, Ltd., Russia): Rb_2CO_3 (99.9%), RE_2O_3 (RE = Eu) (99.99%), and H_3BO_3 (99.99%). Before weighing, Rb_2CO_3 was preheated at 900°C for 24 h to remove any absorbed water. The reagents were weighed on an analytical balance with an accuracy of 0.5 mg. The mixtures of Rb_2CO_3, RE_2O_3, and H_3BO_3 in stoichiometric proportions were thoroughly ground in an agate mortar, slowly heated in a muffle furnace to 500°C at the rate of 1°C/min and held at this temperature for 24 h. Then, the samples were reground and annealed at 700–750°C for 24–72 h until equilibrium is reached. Temperatures were measured with a Pt–PtRh thermocouple. The temperature was controlled to be within ±2°C up to 1200°C with a OMRON controller. The equilibrium was specified when two successive heat treatments result in the identical X-ray patterns.

The powder X-ray diffraction data were recorded by a D8 ADVANCE Bruker AXS diffractometer (Vantec-1 detector) at room temperature using CuK_α radiation and scanning over

the range of 2θ=8-100°. The step size of 2θ was 0.021° and the counting time was 10 s per step. The Rietveld refinement was implemented using TOPAS 4.2 [38]. Almost all peaks were indexed by a trigonal cell (space group *R*32) with the parameters close to those of $K_3YB_6O_{12}$ [33].

Differential scanning calorimetric (DSC) measurements were performed on a STA 449 F1 Jupiter thermoanalyzer (NETZSCH) over the temperature range of 30–900°C in argon flow. Pt crucibles were used as vessels. Pt–PtRh thermocouples were used for measurement. The precision of temperature measurement was ±1°C. The heating and cooling rates were 10°C/min. The second harmonic generation (SHG) response of powder samples was measured with a Q-switched YAG:Nd laser at λ_ω=1064 nm, in the reflection mode. The SHG signal intensities ($I_{2\omega}$) from the sample and from the reference sample (polycrystalline α-SiO_2) were registered comparatively. Infrared (IR) spectra were carried out to specify the boron coordination in the new compounds. The mid-infrared spectra were obtained at room temperature using a Nicolet-380 infrared spectrophotometer with KBr pellets as standards. The spectra were obtained in the range from 500 to 2000 cm^{-1} with the resolution of 1 cm^{-1}.

3. Results and discussions

The synthesized product was a free-flowing powder. Examination of the XRD pattern of the synthesized compound revealed close similarity to that of $K_3YB_6O_{12}$. Therefore, an initial structure model and atomic positions of $K_3YB_6O_{12}$ were adopted for the structure refinement. $Rb_3EuB_6O_{12}$ crystallizes in the noncentrosymmetric space group *R*32, as it was supported by SHG measurements. The intensity of the green light (λ = 532 nm) produced by $Rb_3EuB_6O_{12}$ is about 16 of that produced by the α-quartz standard. The intensity of SHG effect is similar to that of $K_3YB_6O_{12}$ [33]. Weak SHG effect of $Rb_3EuB_6O_{12}$ and $K_3YB_6O_{12}$ is due to the B_5O_{10} groups which are aligned in the almost opposite directions and, consequently, contribute little to the SHG effect.

The refinement results and atomic coordinates obtained for $Rb_3EuB_6O_{12}$ are summarized in Tables 1, 2 and the final Rietveld profiles are shown in Fig. 1. The $Rb_3EuB_6O_{12}$ structure contain a three-dimensional framework composed of $[B_5O_{10}]^{5-}$ groups bridged by Eu-O polyhedra (Fig. 2). The $[B_5O_{10}]^{5-}$ group consists of one BO_4 tetrahedron and four BO_3 triangles that form double B-O rings via the common tetrahedron (Fig. 2). Each $[B_5O_{10}]^{5-}$ group is linked to four different Eu-O polyhedra and likewise each Eu-O polyhedra is connected to four neighboring $[B_5O_{10}]^{5-}$ groups. The Eu-O polyhedra are formed by the face-sharing linked EuO_6 octahedra. Rubidium cations are located in large cavities of the framework structure (Fig. 2). The main bond lengths obtained for $Rb_3EuB_6O_{12}$ are listed in Table 1S, and one can see good agreement of refined values with those reported for the structures containing similar $[B_5O_{10}]^{5-}$ units [12].

The dependence of unit cell volume on the element combination in $M_3REB_6O_{12}$ (M = alkaline metals; RE = Y, rare earth metals) crystal family is shown in Fig. 3 [33-37,39]. It should be noted that all compounds besides $K_3TbB_6O_{12}$ have space group *R*32 and similar cell parameters. The K_3(Tb, Lu, Eu)B_6O_{12} compounds have $R3_2$ space group which is non-standard and, moreover, cell parameter *c* is two times smaller in comparison with other compounds [34-36]. The small cell volume of K_3(Tb, Lu, Eu)B_6O_{12} leads to (Tb, Lu, Eu)/K disordering which seems to be not reasonable. As it appears, these structures should be reinvestigated in future in order to understand the difference. For the analysis shown in Fig. 3, the *c* parameters and unit cell volumes reported for K_3(Tb, Lu, Eu)B_6O_{12} were doubled to uniform the space group *R*32 for $M_3REB_6O_{12}$ crystals. Then, as it is seen from Fig. 3, cell volume increases with IR(M) and IR(RE) increase and the curves for Rb and K are practically parallel. In the first approximation, we can write $V = A\times IR(RE) + B\times IR(M) + C$, where A, B, C are some constants. The least-square method was used to estimate A, B, C using the known data about cell volumes of all known compounds and ions radii of constituted elements [33-37,39]. It was found that general formula $V = 2222.0\times IR(RE) + 1212.6\times IR(M) + 904.7$ Å^3 fits very well the cell volumes of all known compounds, including $Rb_3EuB_6O_{12}$ (Fig. 3). It is interesting that A = 2222.0 is almost two times

bigger than B = 1212.6, and this means that the cell volume variation is dominated by changing of IR(RE). Moreover, the similar analysis for cell parameter variation reveals the relations a = 1.12617×IR(RE) + 1.51834×IR(M) + 10.10377 Å and c = 9.05935×IR(RE) + 0.91584×IR(M) + 20.86787 Å. The equation for c cell parameter showed that M cation is of low influence on c parameter value, but RE provides ten times bigger effect. In the $M_3REB_6O_{12}$ structure, this effect can be associated with the column of three $(RE)O_6$ octahedra linked by faces which goes along c-axis. Probably, increasing of RE ion radii cannot be compensated by other ion moving and this leads to strong c-axis stretching.

Using the diagram shown in Fig. 3, some predictions are possible for discovery of new NCS borates with the $K_3YB_6O_{12}$ structure. The existence of compounds (RE = Yb, Tm, Er, Ho, Dy, Gd) and $Rb_3REB_6O_{12}$ (RE = Pm, Sm) seems to be evident. Besides, the formation of solid solutions $(K,Rb)_3REB_6O_{12}$ may be reasonably supposed for many RE. The situation with M = Na, Cs is less clear. However, the existence of solid solutions $(Na,Rb)_3REB_6O_{12}$ with averaged IR(Na,Rb) close to that of IR(K) is expected with high probability. Above this, the existence of solid solutions $(Na,K)_3REB_6O_{12}$ with partial substitution of Na for K can not be excluded at Na/K ratios when the impact of averaged IR(Na,K) on the structural parameters is not strong.

The synthesized powder $Rb_3EuB_6O_{12}$ possesses light-cream tint, as evident in Fig. 4, that is a common characteristic of Eu-containing oxides [40-42]. Typical SEM image of the particle is shown in Fig. 5. Under the selected technological conditions, synthesis resulted in the agglomerates 5-20 μm in size formed by strongly coalescent individual grains with diameter below 1 μm. This micromorphology is common in oxides when temperature/time conditions used for the solid state reaction are high enough for the active grain interdiffusion [16,34,37,43,44]. It should be pointed that the $Rb_3EuB_6O_{12}$ particles possessed strong charging effect during SEM measurements and this verifies their very low conductivity common for the oxides without oxygen vacancies.

The recorded DSC curve is shown in Fig. 6. Only sharp endothermic peak associated with melting at 1101 K was observed for $Rb_3EuB_6O_{12}$. The powder XRD pattern measured for the sample obtained by the melt solidification is shown in Fig. 1S. Among of the crystalline residues, the presence of $Rb_2(B_4O_5OH)_4)(H_2O)_{3.6}$ (PDF 70-4576), $EuBO_3$ (PDF 74-1931), B_2O_3 (73-1550) was revealed [45-47]. As it is assumed, the $Rb_2(B_4O_5OH)_4)(H_2O)_{3.6}$ compound appeared due to chemical interaction with the air during cooling because of high hygroscopicity of the decomposition products. Besides the crystalline compounds, amorphous component is significant. Thus, it can be concluded that $Rb_3EuB_6O_{12}$ melts incongruently. The thermal parameters available presently for the $M_3REB_6O_{12}$ borates are summarized in Table 3. Regrettably, for $K_3REB_6O_{12}$ compounds, the melting point was earlier measured only for $K_3YB_6O_{12}$. For other $K_3REB_6O_{12}$ borates, however, the melting points can be estimated on the base of temperature used for solid state synthesis. From the comparison of available experimental data, it can be reasonably assumed that melting points of $K_3REB_6O_{12}$ borates are noticeably higher than those of $Rb_3REB_6O_{12}$ borates.

Raman spectrum obtained from $Rb_3EuB_6O_{12}$ is shown in Fig. 7. According to the XRD data, the main structural unit of $Rb_3EuB_6O_{12}$ is so called pentaborate group $[B_5O_{10}]^{5-}$ [48], which consists of one BO_4 tetrahedron and four BO_3 triangles (Fig. 2). The ν_1, ν_3, and ν_4 normal vibration modes of planar BO_3 triangles are Raman-active and ν_2, ν_3, and ν_4 are infrared-active; all four normal vibrations of tetrahedral BO_4 groups are Raman-active, whereas only ν_3 and ν_4 are infrared-active [49]. The total set of vibrational modes in $Rb_3EuB_6O_{12}$ is summarized in Table 4. In Fig. 7, for comparison, the Raman frequencies of pentaborate ion in potassium pentaborate tetrahydrate single crystal are shown by black vertical lines [50].

The antisymmetric stretching of BO_3 triangles of pentaborate group observed in the range of 1300–1700 cm^{-1} in Raman spectra (Fig. 7) and in the range of 1100–1600 cm^{-1} in IR spectra (Fig. 8). Intensity of these bands are greater in Infrared than in Raman spectra [50]. This region could be also covered by the contribution from Eu^{3+} $^5D_1 \rightarrow {}^7F_2$ luminescence [41], however the

presence of similar band in Raman spectrum of $Rb_3NdB_6O_{12}$ indicates its origin both from boron-oxygen network of $Rb_3EuB_6O_{12}$ crystal and from Eu luminescence. The only unexplained lines are narrow doublet at 1290 and 1310 cm^{-1} that is present in Raman spectrum of $Rb_3EuB_6O_{12}$ but absent in $Rb_3NdB_6O_{12}$. Strong Raman bands for boron compounds in this region are not known. Therefore, we must ascribe it to the component of Eu luminescence that experienced red shift by 100 cm^{-1} in the borate under study with respect to europium molybdate. However, this suggestion needs additional study. Symmetric stretching of BO_3 triangles can be found at 875–1000 cm^{-1} in Raman [49, 51] and at 850–975 cm^{-1} in Infrared spectra.

An asymmetric stretching of BO_4 should be observed in Raman spectra in the range of 1000–1150 cm^{-1}, however, intensity of these bands are very weak [50], in the case of Infrared spectra these vibrations are clearly observed at 975–1100 cm^{-1}. Raman bands in the region of 725–800 cm^{-1} are symmetric stretching of BO_4 and strong line at 756 cm^{-1} can be used as an indicator of the pentaborate group presence [52]. Symmetric stretching of BO_4 in the range of 750–810 cm^{-1} can be found in Infrared spectra.

The in-plane bending of BO_3 groups situated in the region of 575–650 cm^{-1} in Raman spectra and at 550–625 cm^{-1} in Infrared spectra. The out-of-plane of BO_3 bending should be inactive in Raman and can be found at 650–750 cm^{-1} in Infrared spectra.

Raman bands in the region of 575–650 cm^{-1} in Raman spectra and spectral region of 550 – 625 cm^{-1} of IR are related to the ν_4 bending modes of BO_4. Strong peak at 530 cm^{-1} is a so-called symmetric pulse vibration of pentaborate anion [53]. Medium bands in the region of 375–500 are ν_2 bending modes of BO_4 groups.

Except for B_5O_{10} vibrations, Raman spectrum of $Rb_3EuB_6O_{12}$ must contain external vibrations involving the movement of heavy constituents (Eu and Rb) of crystal lattice. These vibrations are present in Raman spectrum as the bands observed below 150 cm^{-1}. Rotational and translational modes of $[B_5O_{10}]^{5-}$ ion as the whole are positioned in the same frequency range, bending modes are in the range of 200 – 375 cm^{-1}.

Fig. 9 depicts the luminescence spectrum of $Rb_3EuB_6O_{12}$ excited at 514.5 nm. The infrared part of the spectrum is multiplied by 400 to enhance visibility of longer-wavelength Eu^{3+} luminescent multiplets. All seven $^5D_0$ - ^7F_J bands are present in this spectrum. Noticeable feature of this spectrum is the prominent ultranarrow line corresponding to $^5D_0$ - $^7F_0$ transition. Comparison of this line with the same line in recently studied europium molybdate [41] is presented in Fig. 2S. The position of the ultranarrow line in the $Rb_3EuB_6O_{12}$ borate is at 579.8 nm and its width is 5 cm^{-1} (FWHM), while in α-$Eu_2(MoO_4)_3$ it is at 580.4 nm with the width 2 cm^{-1}. As one can extract from the structural results, Eu ions occupy three inequivalent sites within the unit cell of $Rb_3EuB_6O_{12}$ crystal. Two of them are featured by C_3 local symmetry of Eu^{3+} ion while the third one possesses D_3 local symmetry. The selection rules allow $^5D_0$ - $^7F_0$ transition only in case of symmetry as low as C_3 but not in case of D_3. Therefore, observed ultranarrow line must be ascribed to the first two sites with the required symmetry.

In most crystals and glasses containing Eu^{3+} ions, the hypersensitive $^5D_0$ - $^7F_2$ transition is typically the most intense since its oscillator strength is due to the lack of inversion symmetry, and such environment of Eu^{3+} ion is rather common. This is the case of our borate, too. However, comparing the spectra of $Rb_3EuB_6O_{12}$ and earlier studied α-$Eu_2(MoO_4)_3$, we see that the intensity of ultranarrow line with respect to hypersensitive band is much higher in borate than in molybdate. To reveal the nature of this feature we recorded the luminescence spectra of equal amounts of both materials at the same spectrometer settings, in order to enable approximate comparison of the intensities of lines (difference in the Eu content and in the absorption cross sections was not accounted for). The intensity of hypersensitive band is 43.7 times smaller in borate than in molybdate while the intensity of ultranarrow line is only 2.1 times smaller. Since both emissions originate from the same starting level, we can deduce that breaking of the central inversion of local environment of all three inequivalent Eu sites in $Rb_3EuB_6O_{12}$ leads to smaller oscillator strength gain of hypersensitive transition than the breaking of the mirror symmetry in C_3 sites, with respect to the same symmetry violations in a single C_1 site of Eu in α-$Eu_2(MoO_4)_3$.

Another interesting feature of borate $Rb_3EuB_6O_{12}$ is the shape of hypersensitive $^5D_0$ - $^7F_2$ transition band which is dominated by a single narrow peak at 611 nm, as shown in Fig. 9, This feature is, in principle, favorable for obtaining lasing at this transition since amplification cross section is not smeared over wide spectral region but it is concentrated in the vicinity of peaking wavelength (Fig. 3S). However, due to the above-described fact that concentration quenching in $Rb_3EuB_6O_1$ is stronger than in α-$Eu_2(MoO_4)_3$, the use of lower Eu content must be recommended, e.g. by diluting Eu^{3+} with Gd^{3+} which possesses closest ionic radius.

4. Conclusions

In the $K_3YB_6O_{12}$ borate family, new noncentrosymmetric Rb-containing borates $Rb_3REB_6O_{12}$ (RE = Nd, Eu) are discovered and this greatly extends the nomenclature of the noncentrosymmetric borate crystals available in this family. As it is clear, this borate family covers a wide range of RE elements, including those appropriate for laser and photonic applications. Thus, more wide and deep investigation of the compounds related to the $M_3REB_6O_{12}$ (M = alkaline metals; RE = Y, rare earth metals) crystal family is topical, including search for new compounds and solid solution, development of the crystal growth methods, the observation of piezoelectric and luminescence characteristics. This opens the door for the estimation of the potential of this borate family for practical applications. Luminescence properties of Eu ion in conjunction with nonlinear properties of the matrix indicate that Eu-activated crystals of the $Rb_3REB_6O_{12}$family (here RE = Gd, La and probably Y and Lu) are good candidates for self-doubling Eu-lasing media.

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Table 1. Crystallographic data obtained for $Rb_3EuB_6O_{12}$ phase at room temperature

Compound	$Rb_3EuB_6O_{12}$
Sp. Gr.	*R*32
a, Å	13.4732 (2)
c, Å	30.8424 (6)
V, Å^3	4848.6 (2)
Z	3
2θ-interval, º	5-100
No. of reflections	658
No. of refined parameters	98
R_{wp}, %	2.26
R_p, %	1.63
R_{exp}, %	0.62
χ^2	3.66
R_B, %	1.07

Table 2. Fractional atomic coordinates and isotropic displacement parameters (Å^2) of $Rb_3EuB_6O_{12}$

	x	*y*	*z*	B_{iso}
		$Rb_3EuB_6O_{12}$		
Eu1	1/3	2/3	-0.2182 (7)	0.5 (5)
Eu2	1/3	2/3	-0.7192 (7)	0.5 (5)
Eu3	1/3	2/3	2/3	1.5 (8)
Rb1	0.116 (1)	1/3	5/6	1.5 (7)
Rb2	2/3	0.798 (2)	1/3	1.5 (6)
Rb3	0.187 (1)	0.855 (1)	-0.4146 (6)	1.5 (3)
Rb4	1/3	2/3	-0.580 (1)	1.5 (4)
Rb5	1/3	2/3	1/6	0.9 (11)
B1	0.19 (2)	0.78 (2)	-0.278 (5)	2.0 (9)
B2	0.21 (2)	0.77 (2)	-0.514 (6)	2.0 (9)
B3	0.47 (2)	0.89 (2)	-0.651 (6)	2.0 (9)
B4	0.66 (2)	1.08 (1)	-0.419 (6)	2.0 (9)
B5	0.47 (2)	0.93 (1)	-0.891 (5)	2.0 (9)
O1	0.27 (1)	0.769 (8)	-0.482 (2)	1.5 (5)
O2	0.170 (7)	0.566 (8)	-0.682 (2)	1.5 (5)
O3	0.391 (7)	0.806 (4)	-0.388 (2)	1.5 (5)
O4	0.284 (6)	0.773 (5)	-0.769 (2)	1.5 (5)
O5	0.516 (7)	0.855 (7)	-0.620 (3)	1.5 (5)
O6	0.521 (7)	0.851 (8)	-0.542 (3)	1.5 (5)
O7	0.497 (6)	1.010 (5)	-0.648 (2)	1.5 (5)
O8	0.667 (6)	0.808 (5)	-0.525 (2)	1.5 (5)
O9	0.592 (6)	1.036 (9)	-0.579 (3)	1.5 (5)
O10	0.099 (8)	0.730 (6)	-0.750 (3)	1.5 (5)

Table 3. Thermal parameters obtained for borates with general composition $MREB_6O_{12}$ (M = K, Rb RE = Y, rare earths)

$M_3REB_6O_{12}$	Melting temperature, K	Enthalpy, J/g	Reference
$K_3EuB_6O_{12}$	>1023	-	[34]
$K_3TbB_6O_{12}$	>1023	-	[35]
$K_3YB_6O_{12}$	1214	-	[33]
$K_3LuB_6O_{12}$	>1023	-	[36]
$Rb_3NdB_6O_{12}$	1070	–122.4	[37]
$Rb_3EuB_6O_{12}$	1101	–160.6	This work

Table4. Wyckoff positions of atoms and vibrational modes in $Rb_3EuB_6O_{12}$

Atom	Wyckoff position	Mechanical representation
Rb1	9e	$A_1 + 2A_2 + 3E$
Rb2	9d	$A_1 + 2A_2 + 3E$
Eu1, Eu2, Rb4	6c	$A_1 + A_2 + 2E$
Rb5	3b	$A_2 + E$
Eu3	3a	$A_2 + E$
Rb3, B1–B5, O1–O10	18f	$3A_1 + 3A_2 + 6E$
Total		$53A_1 + 57A_2 + 110E$
Acoustic		$A_2 + E$
Optic		$53A_1 + 56A_2 + 109E$
Raman active modes		$53A_1 + 109E$
IR active modes		$56A_2 + 109E$

Captions

Fig. 1. Measured (red), calculated (black) and differential (blue) diffraction patterns of $Rb_3EuB_6O_{12}$.

Fig. 2. Crystal structure of $Rb_3EuB_6O_{12}$.

Fig. 3. The dependence of unit cell volume V on the element combination in $M_3REB_6O_{12}$ (M = alkaline metals; RE = Y, rare earth metals) crystal family.

Fig. 4. The photo image of $Rb_3EuB_6O_{12}$ powder under illumination by white light.

Fig. 5. SEM pattern of the $Rb_3EuB_6O_{12}$ particle.

Fig. 6. DSC measurements the temperature range of 298 – 813 K.

Fig. 7. Raman spectra of $Rb_3EuB_6O_{12}$ in comparison with band positions of pentaborate ion in potassium pentaborate tetrahydrate single crystal [50].

Fig. 8. IR spectra of $Rb_3REB_6O_{12}$ (RE=Nd, Eu).

Fig. 9. High-resolution PL spectrum for $Rb_3EuB_6O_{12}$ recorded at 514.5 nm excitation with a T64000 spectrometer. Long-wavelength part is multiplied by 400 and shifted up.

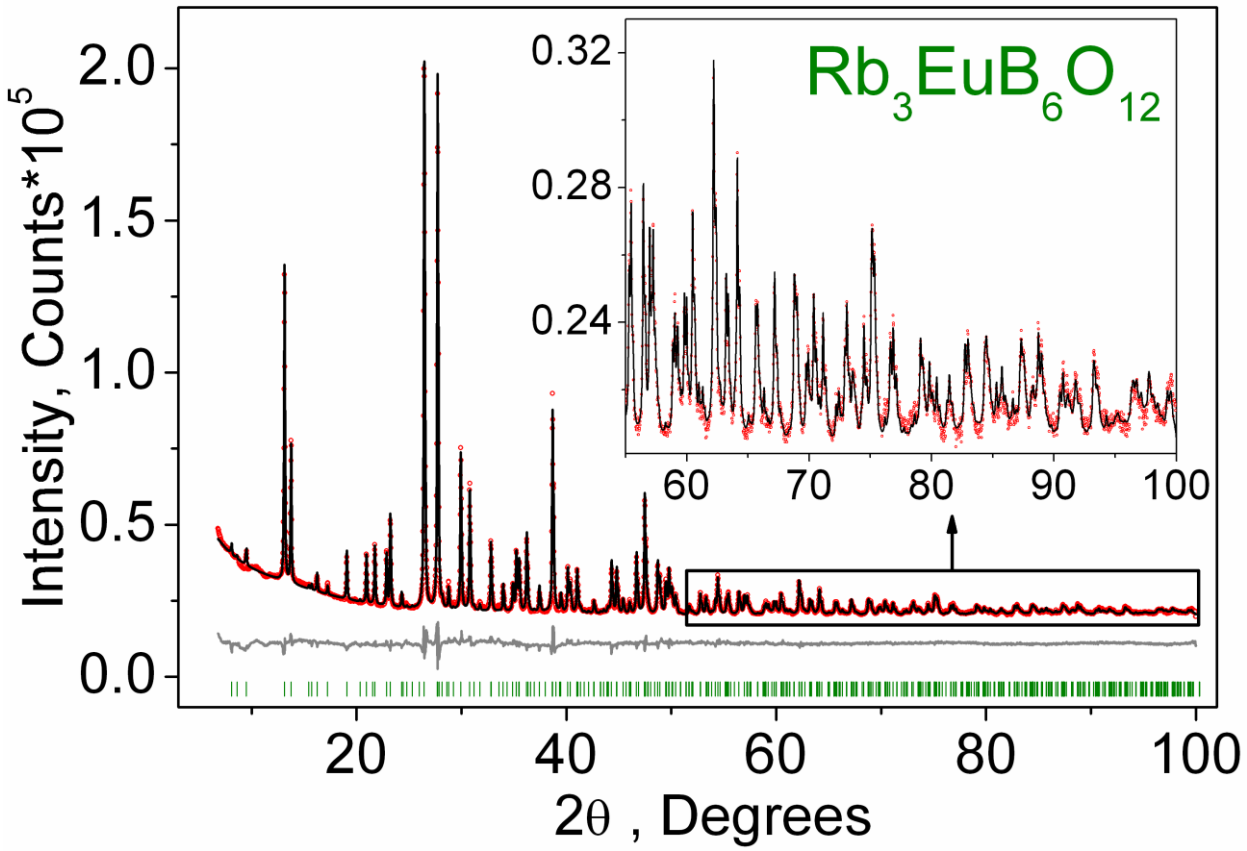

Rb3EuB6O12
Intensity, Counts*10^5
2θ , Degrees
0.0
0.5
1.0
1.5
2.0
20
40
60
80
100
0.24
0.28
0.32
60
70
80
90
100

Fig. 1.

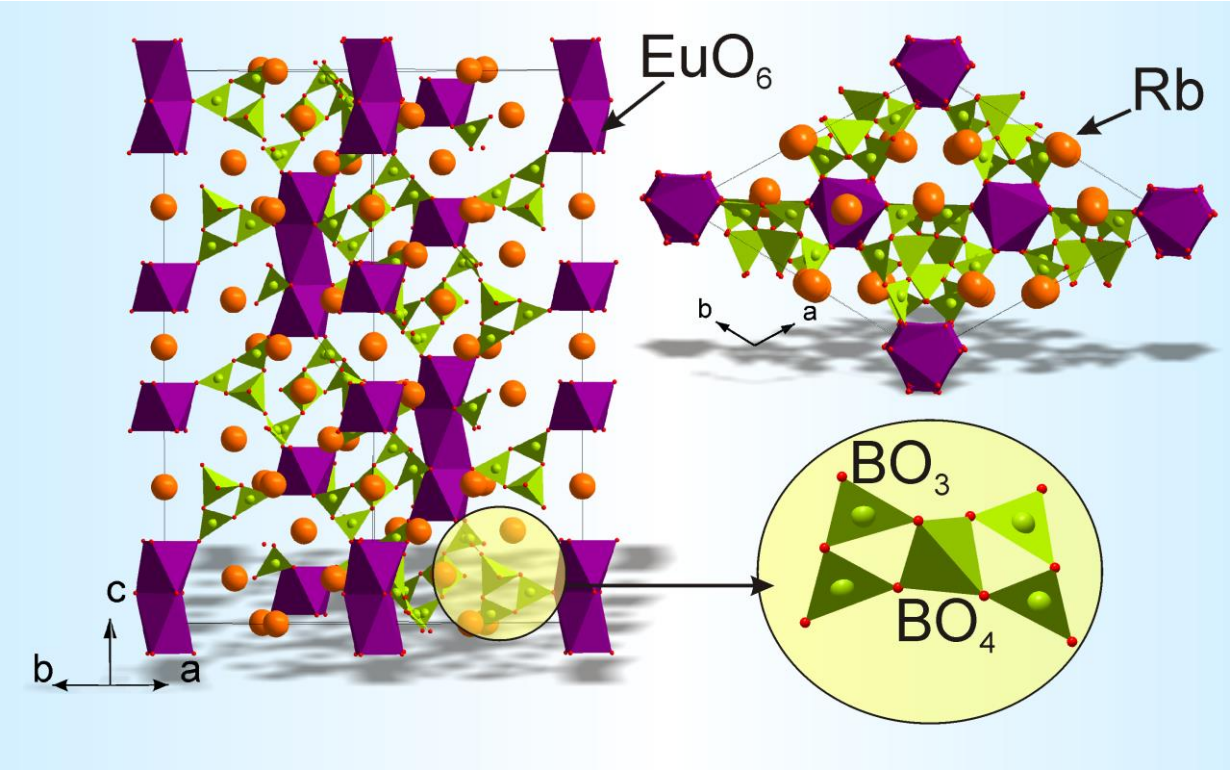


Fig. 2.

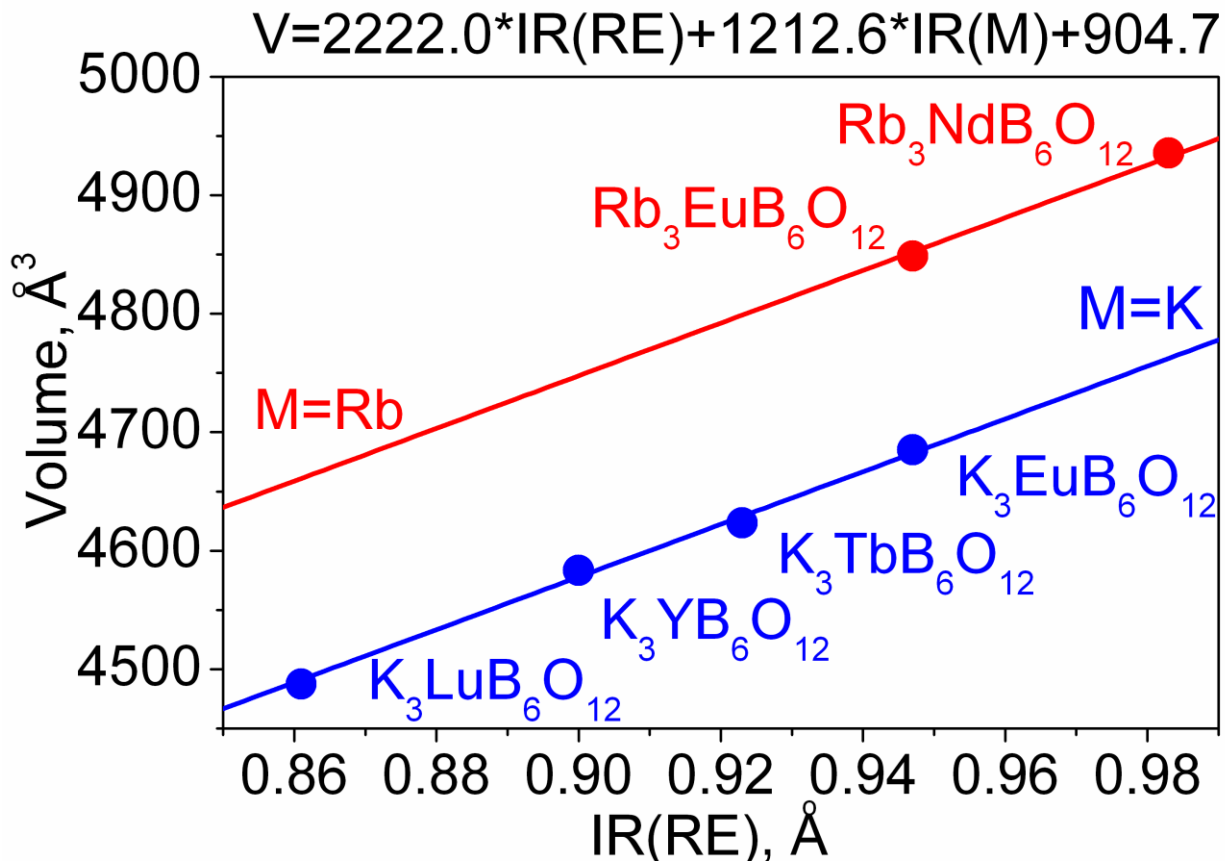

V=2222.0*IR(RE)+1212.6*IR(M)+904.7
Volume, Å3
5000
4900
4800
4700
4600
4500
0.86
0.88
0.90
0.92
0.94
0.96
0.98
IR(RE), Å
Rb3NdB6O12
Rb3EuB6O12
M=K
M=Rb
K3EuB6O12
K3TbB6O12
K3YB6O12
K3LuB6O12

Fig. 3.

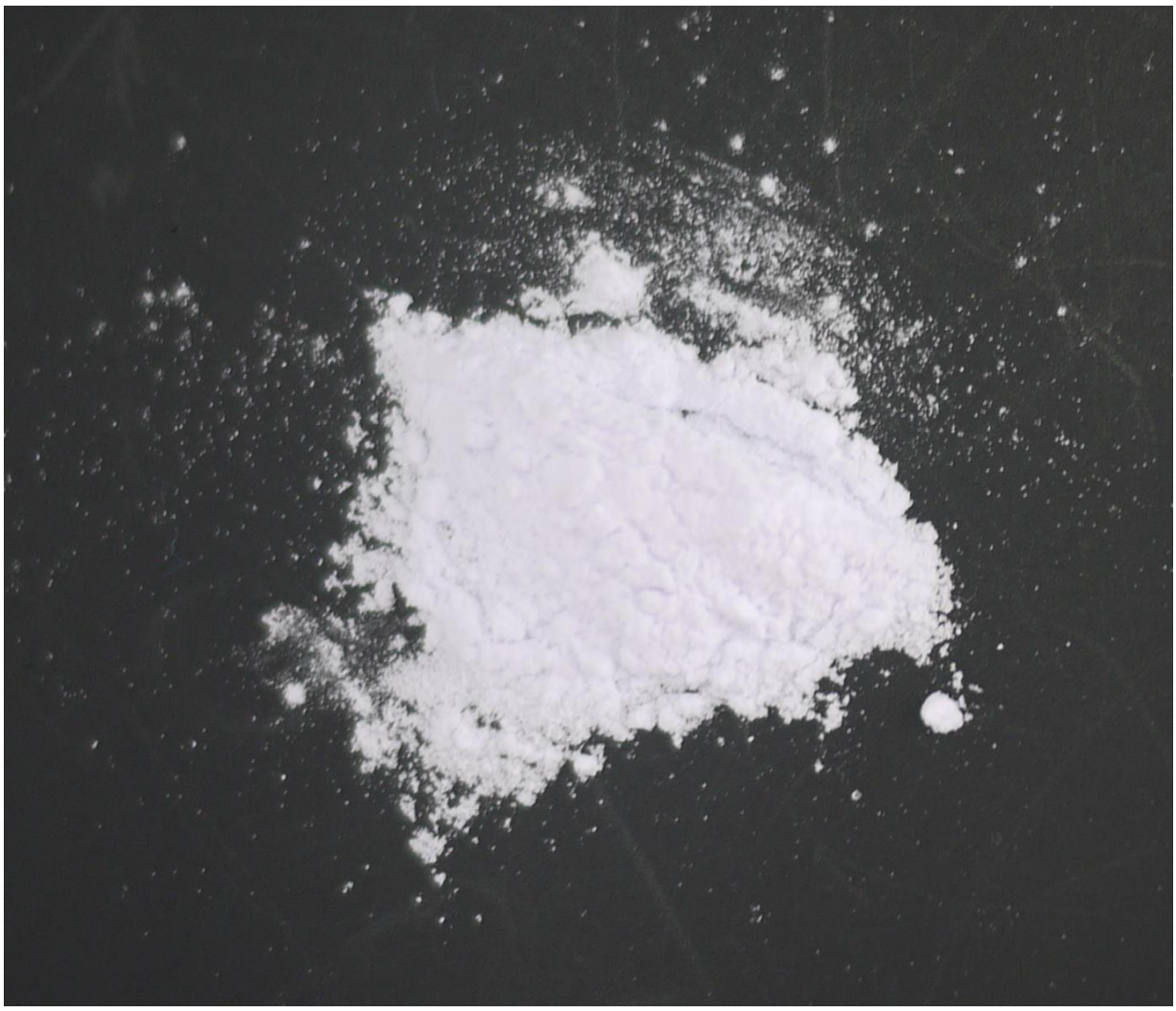
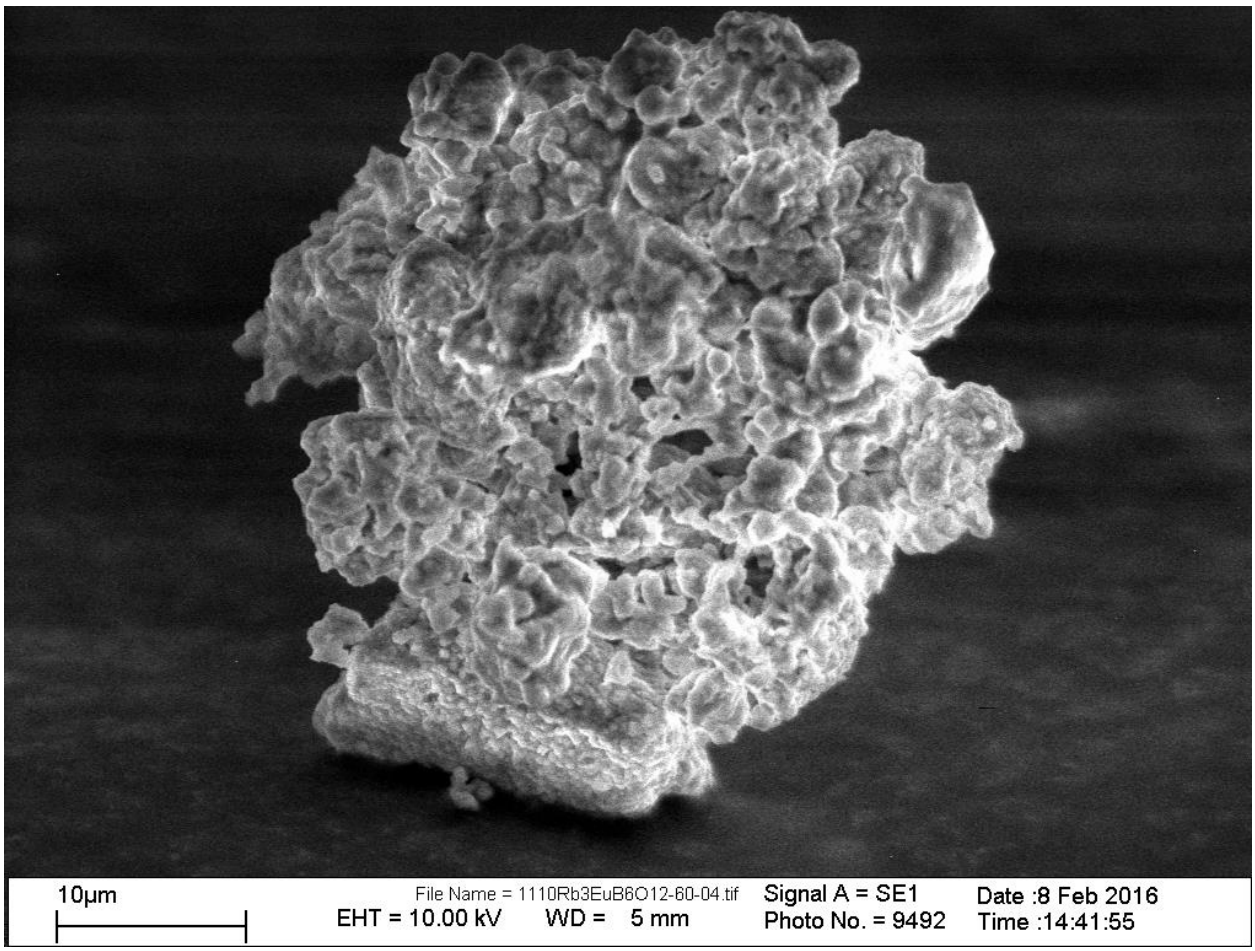


Fig. 4.


10µm
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Signal A = SE1
Photo No. = 9492
Date :8 Feb 2016
Time :14:41:55

Fig. 5.

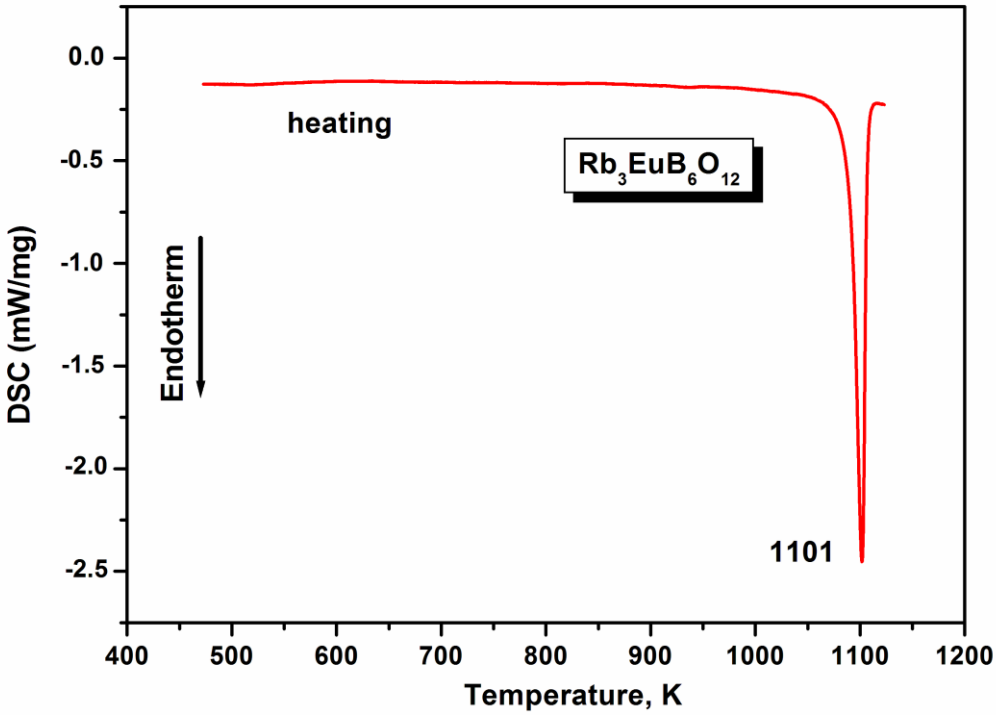
0.0
-0.5
-1.0
-1.5
-2.0
-2.5
DSC (mW/mg)
heating
Rb3EuB6O12
Endotherm
1101
400
500
600
700
800
900
1000
1100
1200
Temperature, K

Fig. 6.

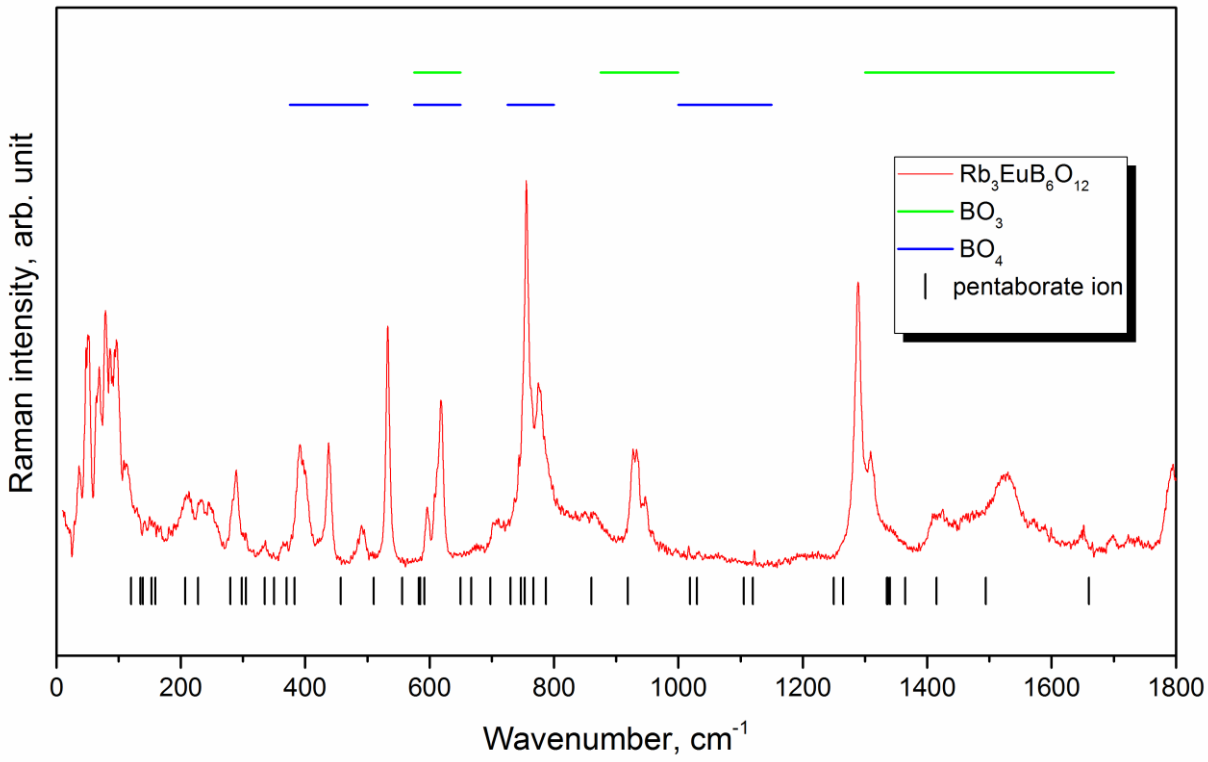
Raman intensity, arb. unit
Wavenumber, cm-1
0
200
400
600
800
1000
1200
1400
1600
1800
Rb3EuB6O12
BO3
BO4
pentaborate ion

Fig. 7.

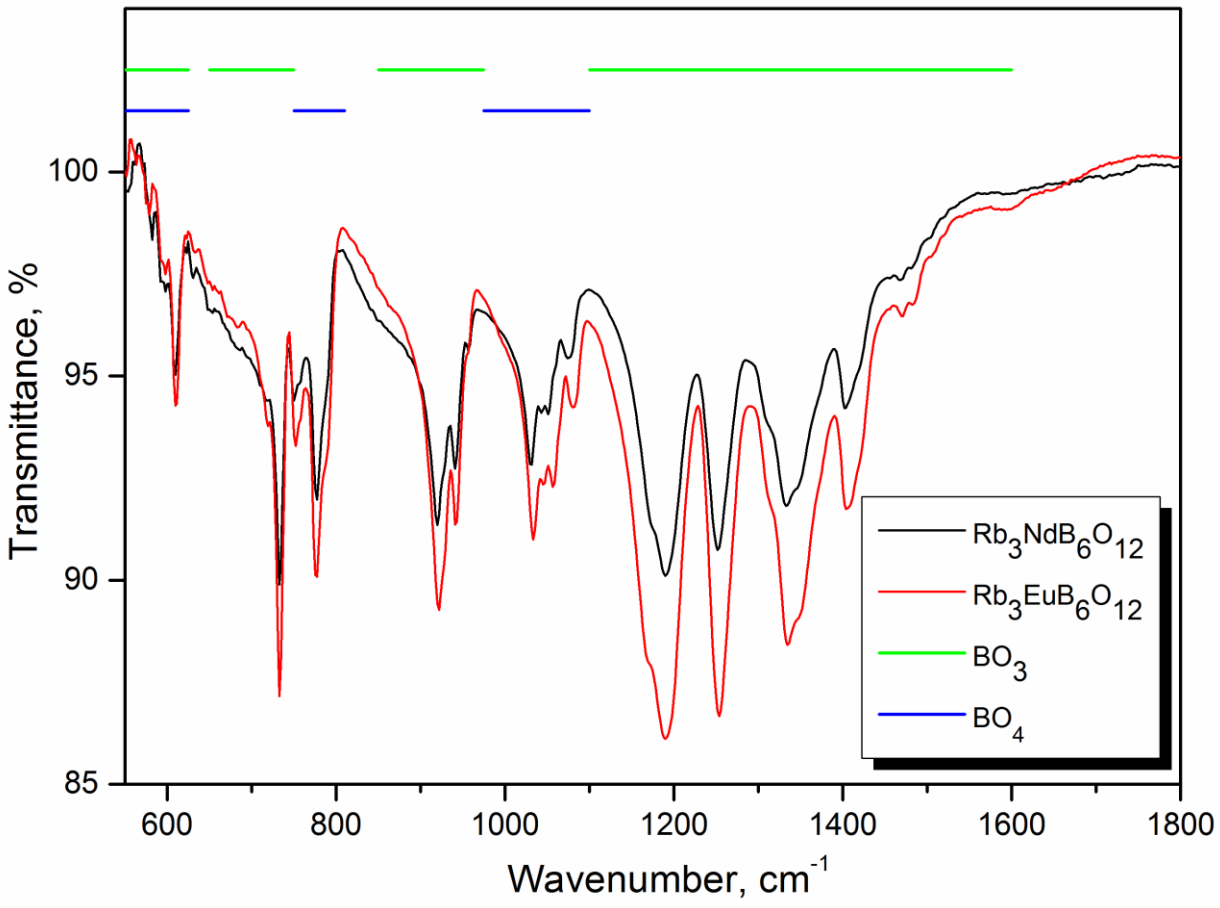
Transmittance, %
100
95
90
85
600
800
1000
1200
1400
1600
1800
Wavenumber, cm-1
Rb3NdB6O12
Rb3EuB6O12
BO3
BO4

Fig. 8.

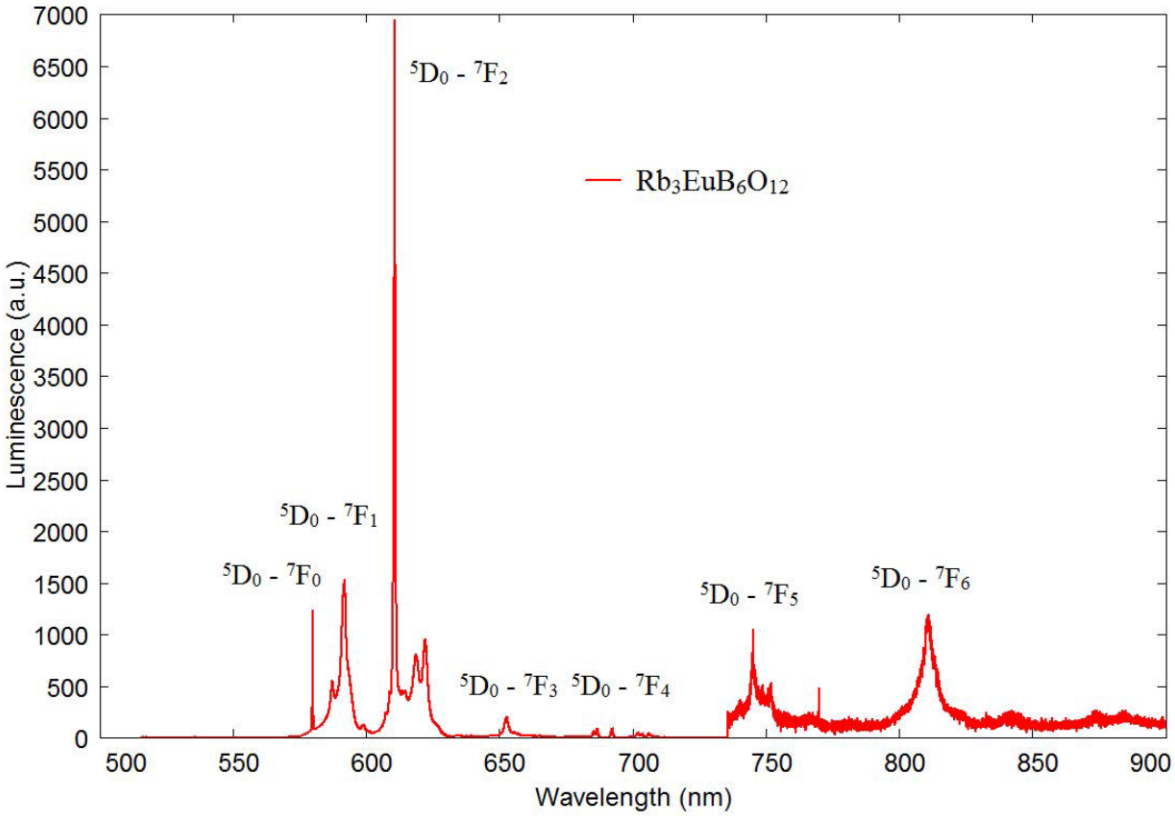
7000
6500
6000
5500
5000
4500
4000
3500
3000
2500
2000
1500
1000
500
0
Luminescence (a.u.)
500
550
600
650
700
750
800
850
900
Wavelength (nm)
$^5D_0$ - $^7F_2$
$Rb_3EuB_6O_{12}$
$^5D_0$ - $^7F_1$
$^5D_0$ - $^7F_0$
$^5D_0$ - $^7F_3$
$^5D_0$ - $^7F_4$
$^5D_0$ - $^7F_5$
$^5D_0$ - $^7F_6$

Fig. 9.

Supporting Information

Structural and spectroscopic properties of new noncentrosymmetric self-activated borate $Rb_3EuB_6O_{12}$ with B_5O_{10} units

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Table 1S. Main bond lengths (Å) in $Rb_3EuB_6O_{12}$

$Rb_3EuB_6O_{12}$			
Eu1—O1^{i}	2.20 (9)	B1—O3^{i}	1.4 (2)
Eu1—O3^{i}	2.48 (6)	B1—O7^{viii}	1.4 (2)
Eu2—O2	2.24 (9)	B1—O9^{viii}	1.3 (2)
Eu2—O4	2.39 (6)	B2—O1	1.3 (2)
Eu3—O3^{ii}	2.36 (5)	B2—O6^{x}	1.4 (2)
Rb1—O1^{iii}	2.88 (9)	B2—O8^{x}	1.4 (2)
Rb1—O6^{iii}	2.93 (7)	B3—O2^{xi}	1.3 (2)
Rb1—O8^{iv}	3.10 (6)	B3—O5	1.4 (2)
Rb1—O9^{iii}	3.00 (7)	B3—O7	1.4 (2)
Rb2—O2^{v}	2.94 (9)	B4—O5^{xii}	1.5 (2)
Rb2—O5^{ii}	2.89 (8)	B4—O6^{xii}	1.5 (2)
Rb2—O10^{vi}	3.07 (9)	B4—O9^{xii}	1.5 (2)
Rb3—O1	2.88 (8)	B4—O10^{xiii}	1.5 (2)
Rb3—O2^{vii}	2.84 (7)	B5—O4^{xiv}	1.3 (2)
Rb3—O3	3.24 (6)	B5—O8^{xv}	1.4 (2)
Rb3—O4^{viii}	2.92 (6)	B5—O10^{xiv}	1.4 (2)
Rb3—O7^{viii}	3.10 (6)		
Rb3—O9^{ix}	3.1 (1)		
Rb3—O10^{vii}	2.95 (8)		
Rb4—O5	2.80 (8)		
Rb4—O6	2.76 (8)		
Rb5—O4^{ii}	2.72 (6)		

Symmetry codes: (i) -*x*+2/3, -*x*+*y*+1/3, -*z*-2/3; (ii) *x*, *y*, *z*+1; (iii) *x*-*y*+2/3, -*y*+4/3, -*z*+1/3; (iv) -*x*+2/3, -*x*+*y*+1/3, -*z*+1/3; (v) -*y*+1, *x*-*y*+1, *z*+1; (vi) -*x*+*y*, -*x*+1, *z*+1; (vii) -*y*+2/3, *x*-*y*+4/3, *z*+1/3; (viii) -*x*+*y*-1/3, -*x*+4/3, *z*+1/3; (ix) *y*-1, *x*, -*z*-1; (x) -*y*+1, *x*-*y*+1, *z*; (xi) *y*-2/3, *x*+2/3, -*z*-4/3; (xii) -*x*+*y*, -*x*+1, *z*; (xiii) *x*-*y*+1, -*y*+2, -*z*-1; (xiv) *x*+2/3, *y*+1/3, *z*+1/3; (xv) -*x*+2/3, -*x*+*y*+1/3, -*z*-5/3; (xvi) -*x*+*y*+1/3, -*x*+5/3, *z*-1/3.

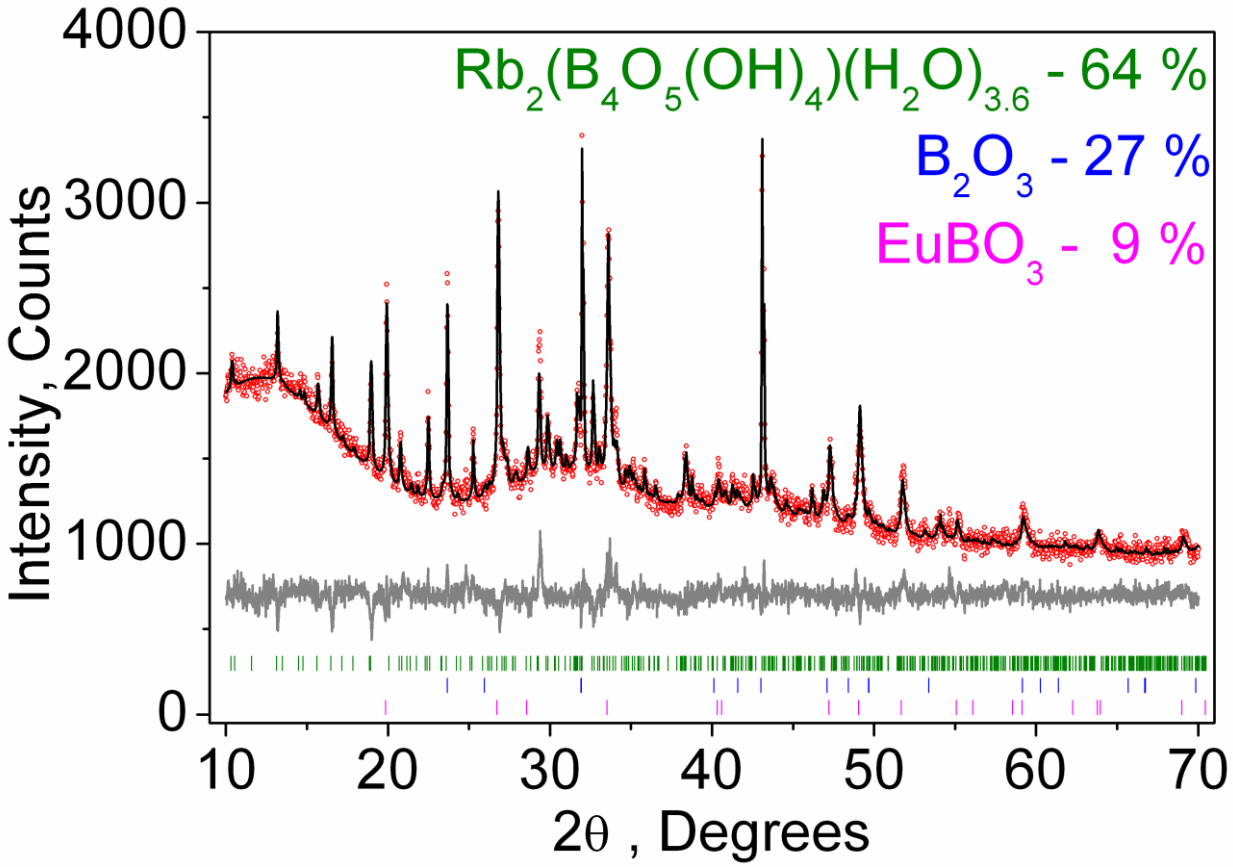

Fig. 1S. Measured (red), calculated (black) and differential (gray) diffraction patterns of the products obtained by thermal decomposition of $Rb_3EuB_6O_{12}$.

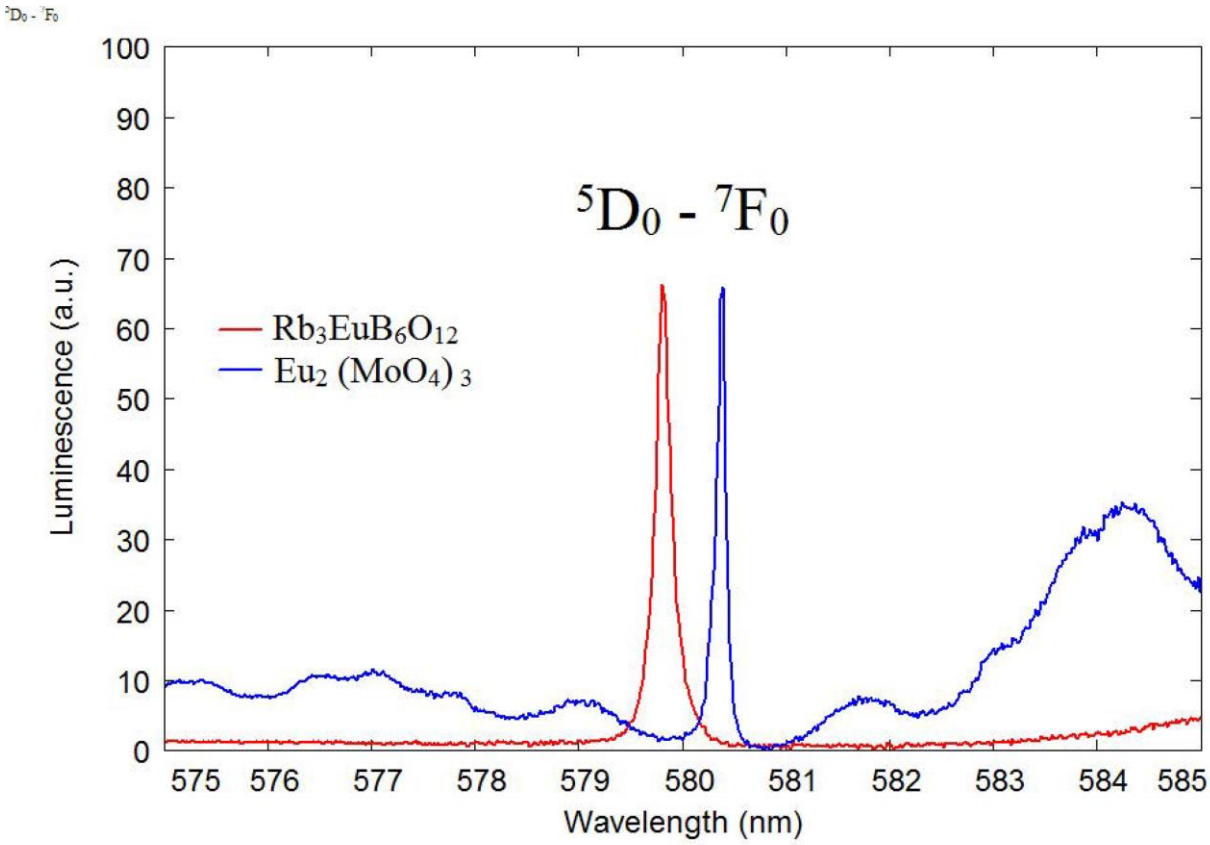

Fig. 2S. High-resolution 514.5 nm excited PL spectra for $Rb_3EuB_6O_{12}$, (red) and $Eu_2(MoO_4)_3$ (blue) in the vicinity of ultranarrow line (T64000 spectrometer). The spectra are equalized at the maximum value.

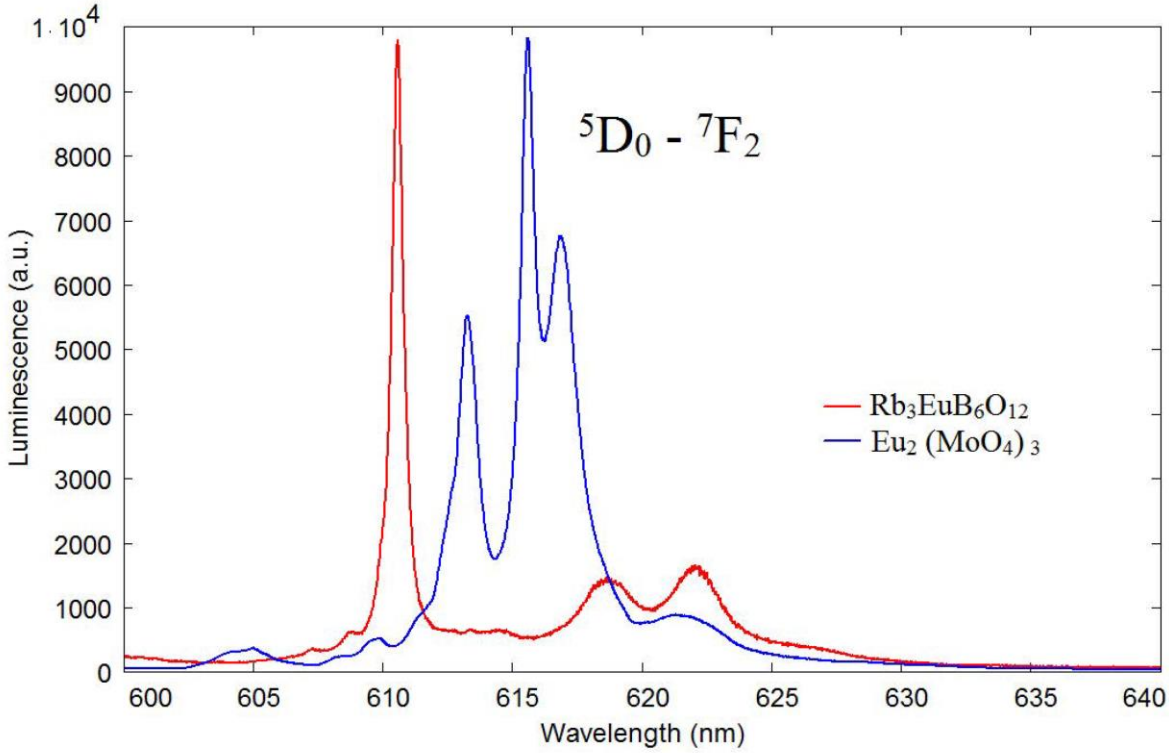

Fig. 3S. Comparison of band shapes in PL spectra for $Rb_3EuB_6O_{12}$, (red) and $Eu_2(MoO_4)_3$ (blue) in the vicinity of hypersensitive transition (T64000 spectrometer). The spectra are equalized at the maximum value.